

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance
 - Volume of an object? ruler or water displacement method
 - Length of an object? Ruler
 - Density of an object? graduated cylinder
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 33 g and has a volume of 3 cm³? 11g/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.
solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of protons and neutrons.

Physical Science Semester Study Guide

Lawrence S. Lerner



Physical Science Semester Study Guide:

A Study Guide for Physics II Gerald E. Buck, 1966 **United States Armed Forces Institute Catalog** United States Armed Forces Institute, 1963 CLEP Official Study Guide College Board, 2001-08 This official guide developed by the sponsors of the CLEP exam includes a CD ROM for Windows that introduces test takers to the new computer based test and provides a tutorial for each exam CLEP Official Study Guide College Entrance Examination Board, 2006-08 A guide to preparing for the thirty four different kinds of college CLEP exams that includes information on deciding which exams to take interpreting the scores taking the tests and other related topics A Student's Guide Through the Great Physics Texts Kerry Kuehn, 2014-09-15 This book provides a chronological introduction to the science of motion and rest based on the reading and analysis of significant portions of Galileo's Dialogues Concerning Two New Sciences Pascal's Treatise on the Equilibrium of Fluids and the Weight of the Mass of Air Newton's Mathematical Principles of Natural Philosophy and Einstein's Relativity Each chapter begins with a short introduction followed by a reading selection Carefully crafted study questions draw out key points in the text and focus the reader's attention on the author's methods analysis and conclusions Numerical and laboratory exercises at the end of each chapter test the reader's ability to understand and apply key concepts from the text Space Time and Motion is the second of four volumes in A Student's Guide through the Great Physics Texts This book grew out of a four semester undergraduate physics curriculum designed to encourage a critical and circumspect approach to natural science while at the same time preparing students for advanced coursework in physics This book is particularly suitable as a college level textbook for students of the natural sciences history or philosophy It also serves as a textbook for advanced high school students or as a thematically organized source book for scholars and motivated lay readers In studying the classic scientific texts included herein the reader will be drawn toward a lifetime of contemplation DOD Pam United States. Office of Armed Forces Information and Education, **Official Study Guide for the CLEP Examinations, 1998** College Entrance Examination Board, 1997 The College Level Examination Program CLEP tests are comprehensive 90 minute exams used to award full college credit for demonstrating college level achievement in a variety of areas in subjects This guide helps students decide which CLEP exams to take helps them interpret their CLEP scores and includes sample questions and answers for all 34 exams and a list of study resources *Catalog of Copyright Entries. Third Series* Library of Congress. Copyright Office, 1976 **Physics for Scientists and Engineers** Lawrence S. Lerner, 1997 This refreshing new text is a friendly companion to help students master the challenging concepts in a standard two or three semester calculus based physics course Dr Lerner carefully develops every concept with detailed explanations while incorporating the mathematical underpinnings of the concepts This juxtaposition enables students to attain a deeper understanding of physical concepts while developing their skill at manipulating equations *Resources in Education* , 1994 **Research in Education** , 1974 *Mathematical Methods for Molecular Science* John E. Straub, Joy Andrews, 2022-08-02 Straub's

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