

Physical Science Study Guide for Nine Weeks Exam

The following chapters will be covered on the exam:

Ch. 1: The Nature of Science

Ch. 15: Classification of Matter

Ch. 16: Solids, Liquids, and Gases

Ch. 17: Properties of Atoms and the Periodic Table

Ch. 18: Radioactivity and Nuclear Reactions

Review questions:

Ch. 1: Nature of Science

1. What does SI stand for? - **International System of Units**

2. List the SI units for mass, length, time, volume, temperature, and density.

Mass = g Length = m Volume = cm^3 or 1 ml Time = sec Temp. = $^{\circ}\text{C}$ Density = g/cm^3

3. Convert the following:

a. 100.0 m to dm = **1000dm**

b. 1.32 mm to Km = **.00000132 Km**

c. 756,000.0 Dm to cm = **756,000,000 cm**

4. What numerical value do the following prefixes represent?

a. Kilo = **1000**

d. deci = **.1**

b. Hecto = **100**

e. centi = **.01**

c. Deka = **10**

f. milli = **.001**

5. How is volume determined for a rock? a cube? **Rock = water displacement**

Cube = $L \times W \times H$

6. How is density calculated? **$D = m/v$**

7. Why is the metric system used in the scientific community? **Easy to use and based on units of 10**

8. What is gathered using the five senses? **Observations**

9. What is an educated guess called? **Hypothesis**

10. Describe the difference in the dependent and independent variable.

Dependent variable = the factor that is measured or observed

Independent variable = the factor that is changed or manipulated

11. When creating a graph, where is each variable placed?

X = independent Y = dependent

12. At the end of an experiment, what is the data compared to?

The control group data

13. What factors are never changed during an experiment? **Control variables**

14. What is an organized manner in which a problem is solved? **Scientific Method**

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Resources in Education, 1998 **Groups and Symmetries** Yvette Kosmann-Schwarzbach, 2009-10-16 Combines material from many areas of mathematics including algebra geometry and analysis so students see connections between these areas Applies material to physics so students appreciate the applications of abstract mathematics Assumes only linear algebra and calculus making an advanced subject accessible to undergraduates Includes 142 exercises many with hints or complete solutions so text may be used in the classroom or for self study *Research in Education*, 1973 Nursing School Entrance Exams Kaplan Nursing, 2016-08-02 Offers a comprehensive review of all tested material on major nursing school entrance assessments including the TEAS HESI PAX RN Kaplan and PSB RN exams Features 2 complete practice tests with detailed answer explanations **Essentials of Physical Chemistry** Don Shillady, 2011-07-27 At a time when U S high school students are producing low scores in mathematics and science on international examinations a thorough grounding in physical chemistry should not be considered optional for science undergraduates Based on the author s thirty years of teaching Essentials of Physical Chemistry merges coverage of calculus with chemistry and molecular physics in a friendly yet thorough manner Reflecting the latest ACS guidelines the book can be used as a one or two semester course and includes special topics suitable for senior projects The book begins with a math and physics review to ensure all students start on the same level and then discusses the basics of thermodynamics and kinetics with mathematics tuned to a level that stretches students abilities It then provides material for an optional second semester course that shows students how to apply their enhanced mathematical skills in a brief historical development of the quantum mechanics of molecules Emphasizing spectroscopy the text is built on a foundation of quantum chemistry and more mathematical detail and examples It contains sample classroom tested exams to gauge how well students know how to use relevant formulas and to display successful understanding of key concepts Coupling the development of mathematical skills with chemistry concepts encourages students to learn mathematical derivations Mini biographies of famous scientists make the presentation more interesting from a people point of view Stating the basic concepts of quantum chemistry in terms of analogies provides a pedagogically useful technique Covering key topics such as the critical point of a van der Waals gas the Michaelis Menten equation and the entropy of mixing this classroom tested text highlights applications across the range of chemistry forensic science pre medical science and chemical engineering In a presentation of fundamental topics held together by clearly established mathematical models the book supplies a quantitative discussion of the merged science of physical chemistry

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