A Company of the Comp	Professional	District

Physical Science — 2nd Semester — Final Exam Study Guide — (2012-2013)

Physical Science Basics

What tooks would you want to use to find the:

Mass of an object? triple beam balance Barrier Strictory No.

Volume of an object? rules or water displacement method Restic St Unit? and or red.

Length of an object? Ruler

Stack: S1 Unit? meter Density of an object? graduated cylinder Blooks St Unit 7 mil.

What is the formula for calculating density? resun/volume.

What is the density of a cube that is 35 g and has a volume of 3 cm 7 Se/cm

Compare the density of solid NLO to liquid NLO, and explain what observable effect this has.

solid HyDis less dense than liquid HyD.

Chemistry Basics.

Dernet (def.) substance made up of one kind of store

Atom (def.) basic building block of matter.

Composed (def.) a majerate that centains atoms or molecules of two or more different elements

4. Molecule (4ef.) a combination of two at more of the same or different atoms that are bonded together

Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created anodels to describe them.

Why have storaic theories channel over time? Better technology come around.

Solumber	Year	Major Discovery/ Contribution	Morne of atomic model
John Deltse	1862	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different mesors.	Billiard Boll Model
J.J. Thoreson	1887	Negatively charged electron	Plum Pudding Model
Niets Bohr	1909	Electrons move fined arbits around the atom's nucleus in specific layers, or siectron shalls. Atoms absorb or give off energy when the electrons move from one shall (grappy level) to enother	Stohr Model
orwest Muldherford	1900	1909 (Based on lifes that positive reflects positive) atom mostly empty space; electrons orbit rendemly around positively charged reciess	Planetary Model

4. Previouslie Testates

- The current periodic table is arranged by increasing atomic number.
- Reackes the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of pretent.
 - The atomic mass represents the number of pentrons and posture

Physical Science Semester Exam Study Guide

Sheila Tobias

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Table of Contents Physical Science Semester Exam Study Guide

- 1. Understanding the eBook Physical Science Semester Exam Study Guide
 - The Rise of Digital Reading Physical Science Semester Exam Study Guide
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Science Semester Exam Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Semester Exam Study Guide
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Science Semester Exam Study Guide
 - Personalized Recommendations
 - Physical Science Semester Exam Study Guide User Reviews and Ratings
 - Physical Science Semester Exam Study Guide and Bestseller Lists

- 5. Accessing Physical Science Semester Exam Study Guide Free and Paid eBooks
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 - Physical Science Semester Exam Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Physical Science Semester Exam Study Guide
 - Interactive Elements Physical Science Semester Exam Study Guide
- 8. Staying Engaged with Physical Science Semester Exam Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Semester Exam Study Guide
- 9. Balancing eBooks and Physical Books Physical Science Semester Exam Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Semester Exam Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Semester Exam Study Guide
 - Setting Reading Goals Physical Science Semester Exam Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Semester Exam Study Guide
 - Fact-Checking eBook Content of Physical Science Semester Exam Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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