

# 16

## Electrostatics

### 16-1 Electrostatic Force

#### Vocabulary

**Electrostatics:** The study of electric charges, forces, and fields.

The symbol for electric charge is the letter “ $q$ ” and the SI unit for charge is the **coulomb (C)**. The coulomb is a very large unit.

$$1 \text{ C} = 6.25 \times 10^{18} \text{ electrons} \quad \text{or} \\ 1 \text{ electron has a charge of } 1.60 \times 10^{-19} \text{ C.}$$

Electrons surrounding the nucleus of an atom carry a negative charge. Protons, found inside the nucleus of the atom, carry a positive charge of  $1.60 \times 10^{-19} \text{ C}$ , while neutrons (which also reside in the nucleus) are neutral. It is important to remember that only electrons are free to move in a substance. Protons and neutrons usually do not move.

When two objects with like charges, positive or negative, are brought near each other, they experience a repulsive force. When objects with opposite charges, one negative and one positive, are brought side by side, they experience an attractive force. These forces can be described with Coulomb’s law.

#### Vocabulary

**Coulomb’s Law:** Two charged objects attract each other with a force that is proportional to the charge on the objects and inversely proportional to the square of the distance between them.

$$F \propto \frac{q_1 q_2}{d^2}$$

This equation looks very similar to Newton’s law of universal gravitation. As before, the sign  $\propto$  means “proportional to.” To make an equation out of this proportionality, insert a quantity called the **electrostatic constant,  $k$** .

$$k = 9.0 \times 10^9 \text{ N} \cdot \text{m}^2/\text{C}^2$$

The magnitude of Coulomb’s law can now be written as an equation.

$$\text{electrostatic force} = \frac{(\text{electrostatic constant})(\text{charge 1})(\text{charge 2})}{(\text{distance})^2} \quad \text{or} \quad F = \frac{kq_1 q_2}{d^2}$$

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