

3.1 Solution

Given: initial gas volume = $5.667 \times 10^{-3} \text{ m}^3$
 initial gas temperature = 294.3 K
 initial gas pressure = 10.13 bar
 gas doubles in volume to $1.133 \times 10^{-2} \text{ m}^3$
 change in volume = $5.667 \times 10^{-3} \text{ m}^3$

Choose, as the system, the universe or surroundings exclusive of the gas in the cylinder. Then, by the constraints given, including no friction in surroundings:

$$Q = 0$$

and

$$-W = P_a (V_f - V_i) + m_p g \Delta Z$$

where

P_a = atmospheric pressure = 1.013 bar
 $V_f - V_i$ = change in system volume = $-5.667 \times 10^{-3} \text{ m}^3$
 m_p = mass of piston = 225 kg
 g = acceleration due to gravity = 9.81 m/s²
 ΔZ = variation in vertical position of piston
 = -0.3048 m if piston rises
 = 0 if piston moves horizontally
 = +0.3048 m if piston falls

Case 1: Piston rises (note 1 bar = 10^5 N/m^2)

$$W = -(1.013 \times 10^5) (-5.667 \times 10^{-3}) - (225) (9.81) (-0.3048) = +1.250 \text{ kJ}$$

$$\Delta E_{\text{univ}} = Q + W = 1.250 \text{ kJ} = -\Delta E_{\text{gas}} = -N_g C_v (T_f - T_i)$$

$$N_g = \text{moles of gas} = \frac{(10.13 \times 10^5) (5.667 \times 10^{-3})}{(8.314) (294.3)} = 2.346$$

$$C_v = 5 \text{ Btu/lb-mol } ^\circ\text{R} = 20.93 \text{ Btu/mol } ^\circ\text{K}$$

Thermodynamics And Its Applications 3rd Edition Solution Manual

Yan Bai



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