

## Thermochemistry Practice Problems (Ch. 6)

1. Consider 2 metals, A and B, each having a mass of 100 g and an initial temperature of 20°C. The specific heat of A is larger than that of B. Under the same heating conditions, which metal would take longer to reach 21°C? Explain your reasoning.
2. A 10.0-g sheet of gold at a temperature of 18°C is placed flat on a 20.0-g sheet of iron at 55.6°C. What is the final temperature of the metals? Assume no heat is lost to the surroundings. ( $c_{Au} = 0.129 \text{ J/g}^\circ\text{C}$ ,  $c_{Fe} = 0.444 \text{ J/g}^\circ\text{C}$ )
3. A 8.6-g piece of aluminum, heated to 100°C, is placed in a coffee cup calorimeter that initially contains 402.4 grams of water at 25°C. If the final temperature is 26.4°C, what is the specific heat of aluminum in  $\text{J/g}^\circ\text{C}$ ?
4. Sulfur (2.56 g) is burned in a bomb calorimeter with excess  $\text{O}_2(\text{g})$ . The temperature increases from 21.25 °C to 26.72 °C. The bomb has a heat capacity of 923 J/K, and the calorimeter contains 815 g of water. Calculate the heat evolved, per mole of  $\text{SO}_2$  formed, in the course of the reaction



5. Assume you mix 100.0 mL of 0.200 M CsOH with 50.0 mL of 0.400 M HCl in a coffee-cup calorimeter. The following reaction occurs:



The temperature of both solutions before mixing was 22.50 °C, and it rises to 24.28 °C after the acid-base reaction. What is the enthalpy of the reaction per mole of CsOH? Assume the densities of the solutions are 1.00 g/mL and the heat capacities of the solutions are 4.2 J/g · K.

6. Consider the following reaction:



What is the value of  $\Delta H$  if (a) the equation is multiplied throughout by 2? (b) the direction of the reaction is reversed? (c) water vapor instead of liquid water is formed?

7. Using standard heats of formation, calculate  $\Delta H$  for the following reactions:
  - a)  $\text{HCl}(\text{g}) \rightarrow \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq})$
  - b)  $2 \text{ NaOH}(\text{s}) + \text{CO}_2(\text{g}) \rightarrow \text{Na}_2\text{CO}_3(\text{s}) + \text{H}_2\text{O}(\text{g})$
  - c)  $\text{C}_2\text{H}_2(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g})$
  - d)  $\text{NO}_2(\text{g}) \rightarrow \text{N}_2\text{O}_4(\text{g})$
8. Given that  $\Delta H$  for the following reaction is -534 kJ, determine the standard heat of formation of hydrazine,  $\text{N}_2\text{H}_4(\text{l})$ .



# Thermochemistry Practice Problem Answers

**ACT**



## Thermochemistry Practice Problem Answers:

**Chemistry: 1,001 Practice Problems For Dummies (+ Free Online Practice)** Heather Hattori, Richard H. Langley, 2014-04-14 Practice makes perfect and helps deepen your understanding of chemistry Every high school requires a course in chemistry and many universities require the course for majors in medicine engineering biology and various other sciences 1001 Chemistry Practice Problems For Dummies provides students of this popular course the chance to practice what they learn in class deepening their understanding of the material and allowing for supplemental explanation of difficult topics 1001 Chemistry Practice Problems For Dummies takes you beyond the instruction and guidance offered in Chemistry For Dummies giving you 1 001 opportunities to practice solving problems from the major topics in chemistry Plus an online component provides you with a collection of chemistry problems presented in multiple choice format to further help you test your skills as you go Gives you a chance to practice and reinforce the skills you learn in chemistry class Helps you refine your understanding of chemistry Practice problems with answer explanations that detail every step of every problem Whether you re studying chemistry at the high school college or graduate level the practice problems in 1001 Chemistry Practice Problems For Dummies range in areas of difficulty and style providing you with the practice help you need to score high at exam time *Chemistry Workbook For Dummies with Online Practice* Chris Hren, Peter J. Mikulecky, 2017-04-17 Take the confusion out of chemistry with hundreds of practice problems Chemistry Workbook For Dummies is your ultimate companion for introductory chemistry at the high school or college level Packed with hundreds of practice problems this workbook gives you the practice you need to internalize the essential concepts that form the foundations of chemistry From matter and molecules to moles and measurements these problems cover the full spectrum of topics you ll see in class and each section includes key concept review and full explanations for every problem to quickly get you on the right track This new third edition includes access to an online test bank where you ll find bonus chapter quizzes to help you test your understanding and pinpoint areas in need of review Whether you re preparing for an exam or seeking a start to finish study aid this workbook is your ticket to acing basic chemistry Chemistry problems can look intimidating it s a whole new language with different rules new symbols and complex concepts The good news is that practice makes perfect and this book provides plenty of it with easy to understand coaching every step of the way Delve deep into the parts of the periodic table Get comfortable with units scientific notation and chemical equations Work with states phases energy and charges Master nomenclature acids bases titrations redox reactions and more Understanding introductory chemistry is critical for your success in all science classes to follow keeping up with the material now makes life much easier down the education road Chemistry Workbook For Dummies gives you the practice you need to succeed Student Solutions Manual Darrell Ebbing, 2005-03-17 Provides solutions to odd numbered Practice Problems General Problems and Cumulative Skills Problems plus answers to Review Questions Chemistry Workbook For Dummies Peter J. Mikulecky, Katherine Brutlag, Michelle Rose

Gilman, Brian Peterson, 2008-08-06 From liquids and solids to acids and bases work chemistry equations and use formulas with ease Got a grasp on the chemistry terms and concepts you need to know but get lost halfway through a problem or worse yet not know where to begin Have no fear this hands on guide helps you solve many types of chemistry problems in a focused step by step manner With problem solving shortcuts and lots of practice exercises you ll build your chemistry skills and improve your performance both in and out of the science lab You ll see how to work with numbers atoms and elements make and remake compounds understand changes in terms of energy make sense of organic chemistry and more 100s of Problems Know where to begin and how to solve the most common chemistry problems Step by step answer sets clearly identify where you went wrong or right with a problem Understand the key exceptions to chemistry rules Use chemistry in practical applications with confidence

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*Survival Guide to General Chemistry* Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

*General Chemistry Workbook* Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

*Materials Thermochemistry* Oswald Kubaschewski, C. B. Alcock, P. J. Spencer, 1993 Materials Thermochemistry the 6th Edition of Metallurgical Thermochemistry aims to demonstrate the central role of thermochemistry in the understanding and designing of materials and materials processes Extensively revised and up dated the 6th Edition of this classic work includes all the latest developments in experimental methods new methods for estimating thermochemical data for both pure and alloy substances new practical applications of thermochemical calculations and up dated tables of critically evaluated thermochemical data for inorganic substances and binary alloy systems The basic principles of chemical

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**AP Chemistry For Dummies** Peter J. Mikulecky, Michelle Rose Gilman, Kate Brutlag, 2008-11-13. A practical and hands on guide for learning the practical science of AP chemistry and preparing for the AP chem exam. Gearing up for the AP Chemistry exam, AP Chemistry For Dummies is packed with all the resources and help you need to do your very best. Focused on the chemistry concepts and problems the College Board wants you to know, this AP Chemistry study guide gives you winning test taking tips, multiple choice strategies, and topic guidelines as well as great advice on optimizing your study time and hitting the top of your game on test day. This user friendly guide helps you prepare without perspiration by developing a pre test plan, organizing your study time, and getting the most out of your AP course. You'll get help understanding atomic structure and bonding, grasping atomic geometry, understanding how colliding particles produce states, and so much more. To provide students with hands on experience, AP chemistry courses include extensive labwork as part of the standard curriculum. This is why the book dedicates a chapter to providing a brief review of common laboratory equipment and

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**Elementary Chemical Reactor Analysis** Rutherford Aris, 2013-09-03 Elementary Chemical Reactor Analysis focuses on the processes reactions methodologies and approaches involved in chemical reactor analysis including stoichiometry adiabatic reactors external mass transfer and thermochemistry The publication first takes a look at stoichiometry and thermochemistry and chemical equilibrium Topics include heat of formation and reaction measurement of quantity and its change by reaction concentration changes with a single reaction rate of generation of heat by reaction and equilibrium of simultaneous and heterogeneous reactions The manuscript then offers information on reaction rates and the progress of reaction in time Discussions focus on systems of first order reactions concurrent reactions of low order general irreversible reaction variation of reaction rate with extent and temperature and heterogeneous reaction rate expressions The book examines the interaction of chemical and physical rate processes continuous flow stirred tank reactor and adiabatic reactors Concerns include multistage adiabatic reactors adiabatic stirred tank stability and control of the steady state mixing in the reactor effective reaction rate expressions and external mass transfer The publication is a dependable reference for readers interested in chemical reactor analysis

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