

Reinforcement Worksheet – Properties of Water

KEY CONCEPT: Water's unique properties allow life to exist on Earth.

The structure of the water molecule gives water unique properties. Water is a polar molecule, which means that it has a region with a slight negative charge (the oxygen atom), and a region with a slight positive charge (the hydrogen atoms). The oppositely charged regions of water molecules interact to form hydrogen bonds. A hydrogen bond is an attraction between a slightly positive hydrogen atom and a slightly negative atom. Hydrogen bonds are responsible for several important properties of water.

- **High specific heat:** Water resists changes in temperature; it must absorb a large amount of heat energy to increase in temperature.
- **Cohesion:** The attraction among molecules of a substance is called cohesion. Cohesion due to hydrogen bonds makes water molecules "stick" together.
- **Adhesion:** The attraction among molecules of different substances is called adhesion. Water molecules "stick" to many other materials because of hydrogen bonds.

Many compounds that are important for life dissolve in water. Water is the largest component of cells' interiors, and chemical reactions in the cell take place in this water. When one substance dissolves in another, a solution is formed. The substance present in the greatest amount is called the solvent. Substances that are present in lower amounts and dissolve in the solvent are called solutes. Polar solvents, such as water, dissolve polar molecules and ions.

When some substances dissolve in water they break up into ions. A compound that releases a hydrogen ion (a proton) when it dissolves in water is an acid. Bases are compounds that remove, or accept, hydrogen ions. A solution's acidity, or its hydrogen ion concentration, is measured on the pH scale. An acid has a low pH (pH below 7) and a high hydrogen ion concentration. A base has a high pH (pH above 7) and a low hydrogen ion concentration. Organisms must maintain a stable pH. Even a small change in pH can disrupt many biological processes.

1. Explain why water is a polar molecule?

2. How do hydrogen bonds form?

3. What are the two parts of a solution?

4. What types of molecules dissolve easily in water? What types do not dissolve easily in water?

5. What does pH mean?

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F Rizvi



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Table of Contents Reinforcement Worksheet Properties Of Water Answers

1. Understanding the eBook Reinforcement Worksheet Properties Of Water Answers
 - The Rise of Digital Reading Reinforcement Worksheet Properties Of Water Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Reinforcement Worksheet Properties Of Water Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reinforcement Worksheet Properties Of Water Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reinforcement Worksheet Properties Of Water Answers

- Personalized Recommendations
 - Reinforcement Worksheet Properties Of Water Answers User Reviews and Ratings
 - Reinforcement Worksheet Properties Of Water Answers and Bestseller Lists
5. Accessing Reinforcement Worksheet Properties Of Water Answers Free and Paid eBooks
 - Reinforcement Worksheet Properties Of Water Answers Public Domain eBooks
 - Reinforcement Worksheet Properties Of Water Answers eBook Subscription Services
 - Reinforcement Worksheet Properties Of Water Answers Budget-Friendly Options
 6. Navigating Reinforcement Worksheet Properties Of Water Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Reinforcement Worksheet Properties Of Water Answers Compatibility with Devices
 - Reinforcement Worksheet Properties Of Water Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reinforcement Worksheet Properties Of Water Answers
 - Highlighting and Note-Taking Reinforcement Worksheet Properties Of Water Answers
 - Interactive Elements Reinforcement Worksheet Properties Of Water Answers
 8. Staying Engaged with Reinforcement Worksheet Properties Of Water Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reinforcement Worksheet Properties Of Water Answers
 9. Balancing eBooks and Physical Books Reinforcement Worksheet Properties Of Water Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reinforcement Worksheet Properties Of Water Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Reinforcement Worksheet Properties Of Water Answers
 - Setting Reading Goals Reinforcement Worksheet Properties Of Water Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Reinforcement Worksheet Properties Of Water Answers

- Fact-Checking eBook Content of Reinforcement Worksheet Properties Of Water Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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