

4) Oxidation-Reduction Reactions

Oxidation - is a loss of electrons
 - is an increase in oxidation number

Reduction - is a gain of electrons
 - is a decrease in oxidation number

Oxidizing Agent - brings about the oxidation of another species
↑ gets reduced

Reducing Agent - brings about the reduction of another species
↑ causes something else to gain electrons by losing electrons itself (Is Oxidized)
↑ gets oxidized

★ OIL RIG (Oxidation Is Loss; Reduction Is Gain)
 LEO-GER (Lose Electrons Oxidation ~ Gain Electrons Reduction)

Memorize common oxidizing pathways (reducing goes backward)

1. Metal $\xrightarrow{\text{ox}}$ metal ion
 $2n \rightarrow 2n^{2+}$
2. Metal $\xrightarrow{\text{ox}}$ metallic ion $\xrightarrow{\text{ox}}$ halate ion
 $\text{Cu} \rightarrow \text{Cu}^+ \rightarrow \text{Cu}^{2+}$
3. Halide ion $\xrightarrow{\text{ox}}$ halogen $\xrightarrow{\text{ox}}$ halite ion $\xrightarrow{\text{ox}}$ halate ion (similar for $\text{NO}_2^- \xrightarrow{\text{ox}} \text{NO}_3^-$)
 $\text{Cl}^- \rightarrow \text{Cl}_2 \rightarrow \text{HOCl} \rightarrow \text{ClO}_2^- \rightarrow \text{ClO}_3^- \rightarrow \text{ClO}_4^-$

Memorize common oxidizing / reducing agents (see bright orange sheet)

Ex. $\text{MnO}_4^- \rightarrow \text{Mn}^{2+}$ (acidified solution)
 $\text{MnO}_4^- \rightarrow \text{MnO}_2$ (basic solution)
 $\text{Cr}_2\text{O}_7^{2-} \rightarrow \text{Cr}^{3+}$ (acidified solution)

===Identify the substances being oxidized and reduced in each of the following.===



MEMORIZE!

Redox Equations With Answers Key

David A. Ucko



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Table of Contents Redox Equations With Answers Key

1. Understanding the eBook Redox Equations With Answers Key
 - The Rise of Digital Reading Redox Equations With Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Redox Equations With Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Redox Equations With Answers Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Redox Equations With Answers Key
 - Personalized Recommendations
 - Redox Equations With Answers Key User Reviews and Ratings
 - Redox Equations With Answers Key and Bestseller Lists
5. Accessing Redox Equations With Answers Key Free and Paid eBooks
 - Redox Equations With Answers Key Public Domain eBooks
 - Redox Equations With Answers Key eBook Subscription Services
 - Redox Equations With Answers Key Budget-Friendly Options
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