

QUANTUM NUMBERS WORKSHEET

1. State the four quantum numbers, then explain the possible values they may have and what they actually represent.

n – Principal Quantum Number: represents the energy level the electron is in, linked to the periods of the periodic. Can be 1 to 7

l – Secondary Quantum Number/Orbital Shape Quantum number: represents the shape of the orbital- s, p, f, d. l is a range of $n-1$.

m_l – Magnetic quantum number: represents the number of orbits possible. M_l is a range of l .

m_s – Spin Quantum number: represents the electron and its spin. Two possibilities $+1/2$, $-1/2$

2. State the number of possible electrons described by the following quantum numbers

- a. $n = 3, l = 0$ **2**
- b. $n = 3, l = 1$ **6**
- c. $n = 3, l = 2, m_l = -1$ **2**
- d. $n = 5, l = 0, m_l = -2, m_s = -1/2$ **Not possible**

3. Give the n and l values for the following orbitals

- a. 1s **$n=1$ $l=0$**
- b. 3s **$n=3$ $l=0$**
- c. 2p **$n=2$ $l=1$**
- d. 4d **$n=4$ $l=2$**
- e. 5f **$n=5$ $l=3$**

4. What are the m_l values for the following types of orbitals?

- a. s **$m_l=0$**
- b. p **$m_l=-1, 0, 1$**
- c. d **$m_l=-2, -1, 0, 1, 2$**
- d. f **$m_l=-3, -2, -1, 0, 1, 2, 3$**

6. How many possible orbitals are there for $n =$

- a. 4 **s-1, p-3, d-5, f-7 = 16 orbitals**
- b. 6 **9 orbitals**

Transparency 14 Periodic Table Answers

**John Olmsted, Greg Williams, Robert C.
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