

## Periodic Trends Worksheet

*Directions: Use your notes to answer the following questions.*

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- |  |                  |
|--|------------------|
| a. atomic radius (excluding noble gases) | <b>Decreases</b> |
| b. first ionization energy               | <b>Increases</b> |
| c. electronegativity                     | <b>Increases</b> |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **increases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

# Trends In The Periodic Table Answer Key

**K. S. Verma**



## **Trends In The Periodic Table Answer Key:**

**Inorganic Chemistry** Mark Weller, Mark T. Weller, Tina Overton, Jonathan Rourke, Fraser Armstrong, 2014 Leading the reader from the fundamental principles of inorganic chemistry right through to cutting edge research at the forefront of the subject Inorganic Chemistry Sixth Edition is the ideal course companion for the duration of a student's degree The authors have drawn upon their extensive teaching and research experience in updating this established text the sixth edition retains the much praised clarity of style and layout from previous editions while offering an enhanced Frontiers section Exciting new applications of inorganic chemistry have been added to this section in particular relating to materials chemistry and medicine This edition also sees a greater use of learning features to provide students with all the support they need for their studies Providing comprehensive coverage of inorganic chemistry while placing it in context this text will enable the reader to fully master this important subject Online Resource Centre For registered adopters of the text Figures marginal structures and tables of data ready to download Test bank For students Answers to self tests and exercises from the book Videos of chemical reactions Tables for group theory Web links Interactive structures and other resources on [www.chemtube3d.com](http://www.chemtube3d.com)

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- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Trends In The Periodic Table Answer Key. The first chapter will explore what Trends In The Periodic Table Answer Key is, why Trends In The Periodic Table Answer Key is vital, and how to effectively learn about Trends In The Periodic Table Answer Key.

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## **Table of Contents Trends In The Periodic Table Answer Key**

1. Understanding the eBook Trends In The Periodic Table Answer Key
  - The Rise of Digital Reading Trends In The Periodic Table Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Trends In The Periodic Table Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Trends In The Periodic Table Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Trends In The Periodic Table Answer Key
  - Personalized Recommendations
  - Trends In The Periodic Table Answer Key User Reviews and Ratings
  - Trends In The Periodic Table Answer Key and Bestseller Lists
5. Accessing Trends In The Periodic Table Answer Key Free and Paid eBooks
  - Trends In The Periodic Table Answer Key Public Domain eBooks
  - Trends In The Periodic Table Answer Key eBook Subscription Services
  - Trends In The Periodic Table Answer Key Budget-Friendly Options
6. Navigating Trends In The Periodic Table Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Trends In The Periodic Table Answer Key Compatibility with Devices
  - Trends In The Periodic Table Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Trends In The Periodic Table Answer Key
  - Highlighting and Note-Taking Trends In The Periodic Table Answer Key
  - Interactive Elements Trends In The Periodic Table Answer Key

8. Staying Engaged with Trends In The Periodic Table Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Trends In The Periodic Table Answer Key
9. Balancing eBooks and Physical Books Trends In The Periodic Table Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Trends In The Periodic Table Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Trends In The Periodic Table Answer Key
  - Setting Reading Goals Trends In The Periodic Table Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Trends In The Periodic Table Answer Key
  - Fact-Checking eBook Content of Trends In The Periodic Table Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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