#### **GRAPHING PERIODIC TRENDS**

The Periodic Table is arranged according to the Periodic Law. The Periodic Law states that when elements are arranged in order of increasing atomic number, their physical and chemical properties show a periodic pattern.

Define the following terms(starting on page 170):

Atomic radius-

First ionization energy-

Electronegativity-

#### Graph 1 - Atomic Radius as a function of Atomic Number

- A. Create a graph of the atomic radius as a function of atomic number. Plot atomic number on the X axis and atomic radius on the Y axis. Remember to label the axes!
- B. Use a colored pen, pencil or highlighter to trace over the element's period (horizontal row on the periodic table). For example: use GREEN to trace for all of the elements in row 1, then use YELLOW to trace for all of the elements in row 2, then use ORANGE to trace for all the elements in row 3, then use BLUE to trace for all the elements in row 4.

Symbol	Atomic Number	Atomic Radius(pm)					
H	1	37					
He	2	50					
Li	3	123					
Be	4	89					
В	5	80					
C	6	77					
N	7	70					
0	8	66					
F.	9	64					
Ne	10	67					
Na	1.1	157					
Mg	12	136					
AI	13	125					
Si	1.4	1.17					
P	1.5	110					
S	16	104					
CI	17	99					
Air	18	98					
K	19	203					
Ca	20	174					

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- Describe the trend in atomic radius as the atomic number increases across a period.
- 2. Describe the trend in atomic radius as the atomic number increases down a group.

# **Trends In The Periodic Table Graphing Worksheet Answers**

**George Graybill** 

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#### **Table of Contents Trends In The Periodic Table Graphing Worksheet Answers**

- 1. Understanding the eBook Trends In The Periodic Table Graphing Worksheet Answers
  - The Rise of Digital Reading Trends In The Periodic Table Graphing Worksheet Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Trends In The Periodic Table Graphing Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Trends In The Periodic Table Graphing Worksheet Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Trends In The Periodic Table Graphing Worksheet Answers

- Personalized Recommendations
- Trends In The Periodic Table Graphing Worksheet Answers User Reviews and Ratings
- Trends In The Periodic Table Graphing Worksheet Answers and Bestseller Lists
- 5. Accessing Trends In The Periodic Table Graphing Worksheet Answers Free and Paid eBooks
  - Trends In The Periodic Table Graphing Worksheet Answers Public Domain eBooks
  - Trends In The Periodic Table Graphing Worksheet Answers eBook Subscription Services
  - Trends In The Periodic Table Graphing Worksheet Answers Budget-Friendly Options
- 6. Navigating Trends In The Periodic Table Graphing Worksheet Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Trends In The Periodic Table Graphing Worksheet Answers Compatibility with Devices
  - Trends In The Periodic Table Graphing Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Trends In The Periodic Table Graphing Worksheet Answers
  - Highlighting and Note-Taking Trends In The Periodic Table Graphing Worksheet Answers
  - Interactive Elements Trends In The Periodic Table Graphing Worksheet Answers
- 8. Staying Engaged with Trends In The Periodic Table Graphing Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Trends In The Periodic Table Graphing Worksheet Answers
- 9. Balancing eBooks and Physical Books Trends In The Periodic Table Graphing Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Trends In The Periodic Table Graphing Worksheet Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Trends In The Periodic Table Graphing Worksheet Answers
  - Setting Reading Goals Trends In The Periodic Table Graphing Worksheet Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Trends In The Periodic Table Graphing Worksheet Answers

- Fact-Checking eBook Content of Trends In The Periodic Table Graphing Worksheet Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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