

AP Chemistry Unit 5 Review

Thermochemistry

Honors Lecture Notes: Unit 4

Terminology

Temperature measures the average kinetic energy of molecules. Thermal energy measures the total amount of energy in a substance. Heat measures the flow of energy from substances at high temperatures to those at low temps.

Specific heat describes the amount of energy needed to change the temperature of an object. The lower the specific heat of an object is, the easier it is to change the temperature of that object. For two substances with identical compositions, the darker the color the lower the specific heat of that object.

Calorimetry

The quantitative study of how heat affects temperature change is called calorimetry. The most important equation here is:

$$q = mc\Delta T$$

q is heat (in J)

m is mass (in g)

c is specific heat (in J/g°C) T is temperature change (in °C)

If an object is losing energy, the sign for q is negative. If an object is gaining energy, the sign is positive. In a closed system, the amount of heat energy is conserved. For objects to gain energy, other objects must lose that energy first.

Enthalpy of Formation

Enthalpy of formation (ΔH_f) for a substance describes the amount of energy that is absorbed or released when one mole of the substance is created from elements in their standard states. The "standard state" for any element is a single atom except for diatomics (H_2 , O_2 , F_2 , Br_2 , I_2 , N_2 , Cl_2).

Enthalpies of formation are often negative, as the only time you need to break bonds to form a new compound is when creating compounds that have elements which are diatomic in their standard state. Enthalpies of formation are the most common place to see fractional coefficients in front of diatomics to reflect the creation of one mole of a product.



Bond Enthalpies

Breaking bonds requires energy, and the creation of bonds releases energy. When considering how many bonds are broken or formed during a chemical reaction, you must draw the Lewis diagram for all species first. Then, consider how many of those bonds are in each molecule, as well as the coefficient for each molecule.



In the above reaction, four C-H bonds are broken, as are two O=O bonds. 2 C=O bonds are created, as are 4 O-H bonds.

Calculating Enthalpy of Reaction

There are four primary ways to calculate the enthalpy change for a reaction (ΔH_{rxn}).

1) **Calorimetry**- Many chemical reactions are aqueous, and by tracking the temperature of the water you can determine if the reaction itself absorbed or emitted heat. If the water temperature increases, the reaction must have emitted heat (opposite sign). If you divide the heat of reaction by the moles of limiting reactant, you get ΔH_{rxn} .

2) **Enthalpies of Formation**- If you subtract the combined enthalpies of formation of the reactants from the combined enthalpies of formation of the products, you will get the overall enthalpy of reaction. You must modify each species by any coefficient it may have in the balanced equation. (**Products – Reactants**)

3) **Bond Enthalpies**- Subtract the total energy of the bonds formed (products) from the total energy of the bonds broken (reactants) to get the enthalpy of reaction. (**Reactants – Products**)

4) **Hess's Law**- If you know the enthalpy of reactions which, when combined, create a new reaction, you can combine the enthalpies of reaction for the component reactions to determine the enthalpy of reaction for the combined reaction.

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Richard Lloyd Whiteley



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Relativistic Quantum Chemistry Markus Reiher, Alexander Wolf, 2009-02-17 Written by two researchers in the field this book is a reference to explain the principles and fundamentals in a self contained complete and consistent way Much attention is paid to the didactical value with the chapters interconnected and based on each other From beginning to end the authors deduce all the concepts and rules such that readers are able to understand the fundamentals and principles behind the theory Essential reading for theoretical chemists and physicists Book Jacket *Russian Journal of Physical Chemistry*, 1983-07

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