

Unit 7 Study Guide

7.1 – Temperature, Thermal Energy, and Heat

1. What is temperature?

The measure of the average kinetic energy of the particles that make up an object.

2. How are thermal energy and temperature related?

Thermal energy is the sum of the kinetic energy (temperature) and potential energy of particles.

Increased temperature = increased thermal energy

3. What is the difference between thermal energy and heat?

Heat is the energy that is transferred due to temperature differences.

Thermal energy is one type of energy that objects have.

4. How can you calculate changes in thermal energy?

$$Q = m(T_f - T_i)C$$

5. Describe how the motions of particles that make up an object change when the object's temperature increases.

Increasing the temperature means increasing the motion of particles.

6. Describe the energy transfer when you touch a block of ice with your hand.

Energy from your hand transfers to the ice, increasing the energy of the ice.

7. When one object heats another, does the temperature increase of one object always equal the temperature decrease of the other object? Explain.

Only if the 2 objects have the same mass and same specific heat (meaning they are the same material).

Otherwise, no.

8. Explain why water is often used as a coolant.

Because it has a relatively high specific heat (it takes a lot of energy to heat it up so it will absorb a lot of energy before heating up).

9. Estimate the change in thermal energy of water in a pond with a mass of 1000 kg and a specific heat of 4200 J/kg°C if the water cools by 1°C.

$$Q = 1000(1)(4200) = 4,200,000 \text{ J}$$

$$Q = m \Delta T C$$

Transfer Of Energy Study Guide

Neil A. Campbell



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