

Chapter 9 Review

- 1) What does stoichiometry mean? "measuring elements"
- 2) Before performing a stoichiometry problem, you must first do what with the chemical equation? **balance it**
- 3) The coefficients in a chemical equation represent what? **moles**
- 4) In a chemical equation, what relationships are shown? **molar ratios**
- 5) Solve the following mole-mole conversions using this equation:



- a) If 3.5 moles of sodium bicarbonate are decomposed, how many moles of carbon dioxide are produced?

$$3.5 \text{ mol NaHCO}_3 \times \frac{1 \text{ mol CO}_2}{2 \text{ mol NaHCO}_3} = 1.8 \text{ mol CO}_2$$

- b) If 1.29 moles of water are produced by the decomposition of sodium bicarbonate, how many moles of sodium carbonate are produced?

$$1.29 \text{ mol H}_2\text{O} \times \frac{1 \text{ mol Na}_2\text{CO}_3}{1 \text{ mol H}_2\text{O}} = 1.29 \text{ mol Na}_2\text{CO}_3$$

- c) How many moles of sodium bicarbonate are required to produce 19.5 moles of carbon dioxide?

$$19.5 \text{ mol CO}_2 \times \frac{2 \text{ mol NaHCO}_3}{1 \text{ mol CO}_2} = 39.0 \text{ mol NaHCO}_3$$

- 6) Use the equation in # 5 to solve the following mass-mass conversions.

- a) If given 10.5 grams of sodium bicarbonate, how many grams of sodium carbonate can be produced?

$$10.5 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol Na}_2\text{CO}_3}{2 \text{ mol NaHCO}_3} \times \frac{105.99 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} = 6.62 \text{ g Na}_2\text{CO}_3$$

- b) How many grams of sodium bicarbonate would be required to produce 9.0 grams of water?

$$9.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{2 \text{ mol NaHCO}_3}{1 \text{ mol H}_2\text{O}} \times \frac{84.01 \text{ g NaHCO}_3}{1 \text{ mol NaHCO}_3} = 84 \text{ g NaHCO}_3$$

- 7) What conversion factor (mole ratio) would you use to solve for moles of sulfur formed from sulfur dioxide in the reaction



Trends In Size Chemistry Packet Answers

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Computational Chemistry: Reviews Of Current Trends, Vol. 5 Ermanno Gianinetti, Jozek S Kwiatkowski, Jerzy Leszczynski, Piotr Piecuch, Jiri Sponer, Thanh N Truong, Charles A Weatherford, 2000-10-09 This volume comprises six chapters which explore the development and applications of the methods of computational chemistry The first chapter is on new developments in coupled cluster CC theory The homotopy method is used to obtain complete sets of solutions of nonlinear CC equations The correspondence between multiple solutions to the CCSD CCSDT and full CI equations is established and the applications of the new approach in modeling molecular systems are discussed The second chapter reviews the computational theory for the time dependent calculations of a solution to the Schrödinger equation for two electrons and focuses on the development of propagators to the solution The next chapter features a discussion on a new self consistent field for molecular interactions SCF MI scheme for modifying Roothaan equations in order to avoid basis set superposition errors BSSE This method is especially suitable for computations of intermolecular interactions Details of the theory along with examples of applications to nucleic acid base pair complexes are given This chapter is well complemented by the following chapter which reports the current status of computational studies of aromatic stacking and hydrogen bonding interactions among nucleic acid bases The next chapter reveals the possibility of calculating the kinetics of chemical reactions in biological systems from the first principles The last chapter reviews the results of rigorous ab initio studies of the series of derivatives of methane silane and germane The presented molecular and vibrational parameters complement experimental data for these systems In addition the theoretical approach allows the prediction of the effects of halogeno substitutions on their structures and properties

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