

POGIL: Periodic Table Trends

(Adapted from Rush Henrietta CSD)

Why:

The Periodic Table is one of the greatest inventions in the history of man. It allows scientists to predict physical and chemical properties of the elements. Dimitri Mendeleev (a Russian scientist) and Robert Mosley (a British chemist) put together this table in the late 1800's based on properties of the elements known at the time. The trends in these properties as you go across periods and down groups is the subject of this Chem POGIL.

Periodic Table of the Elements

Legend:

- 1. alkali metals
- 2. alkaline earth metals
- 3. transition metals
- 4. metalloids
- 5. nonmetals
- 6. noble gases
- 7. lanthanide series
- 8. actinide series

The table shows elements from Hydrogen (H) to Oganesson (Og). The lanthanide and actinide series are shown below the main table.

Success Criteria:

- Understand the meaning of *atomic radius, reactivity, electronegativity, ionization energy*
- Recognize trends in atomic radius, reactivity, ionization energy and electronegativity as you go across periods and down groups.

Resources:

Reference Table S, Periodic Table, Unit 3 (Chapter 6) Notes, & Chapter 6 in textbook

Prerequisites:

Answer the following questions using the resources above.

- What happens to the number of valence electrons as you go down a group on the periodic table?
- What happens to the number of valence electrons as you go from left to right across a period?
- What happens to the number of energy levels as you go down a group on the periodic table?
- What happens to the number of energy levels as you go from left to right across a period?
- Draw a Lewis structure for a Cl atom.
- Draw a Bohr diagram for the Cl atom. (Draw above!)
- In this lab we will be talking about atomic radius.
 - Look up the atomic radius of Cl on Table S: _____
 - Why is it possible to show the atomic radius using a Bohr diagram but not by using a Lewis structure?
 - On your Bohr diagram draw an arrow to represent the radius of the Cl atom.
- Define *periodic*:
- Define *trend*:

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Peipei Pang



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Table of Contents U3 Pogil Periodic Table

1. Understanding the eBook U3 Pogil Periodic Table
 - The Rise of Digital Reading U3 Pogil Periodic Table
 - Advantages of eBooks Over Traditional Books
2. Identifying U3 Pogil Periodic Table
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an U3 Pogil Periodic Table
 - User-Friendly Interface
4. Exploring eBook Recommendations from U3 Pogil Periodic Table
 - Personalized Recommendations
 - U3 Pogil Periodic Table User Reviews and Ratings
 - U3 Pogil Periodic Table and Bestseller Lists
5. Accessing U3 Pogil Periodic Table Free and Paid eBooks
 - U3 Pogil Periodic Table Public Domain eBooks
 - U3 Pogil Periodic Table eBook Subscription Services
 - U3 Pogil Periodic Table Budget-Friendly Options

6. Navigating U3 Pogil Periodic Table eBook Formats
 - ePub, PDF, MOBI, and More
 - U3 Pogil Periodic Table Compatibility with Devices
 - U3 Pogil Periodic Table Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of U3 Pogil Periodic Table
 - Highlighting and Note-Taking U3 Pogil Periodic Table
 - Interactive Elements U3 Pogil Periodic Table
8. Staying Engaged with U3 Pogil Periodic Table
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers U3 Pogil Periodic Table
9. Balancing eBooks and Physical Books U3 Pogil Periodic Table
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection U3 Pogil Periodic Table
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine U3 Pogil Periodic Table
 - Setting Reading Goals U3 Pogil Periodic Table
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of U3 Pogil Periodic Table
 - Fact-Checking eBook Content of U3 Pogil Periodic Table
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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