

(b) Use Eq. 1.5-12 to get: (Here $T_b = 373.2^\circ\text{K}$)

$$\text{At } 273.2^\circ\text{K} \quad \mu = (2.21 \times 10^{-4}) \exp (3.8)(373.2)/(273.2) = 4.0 \times 10^{-2} \text{ g cm}^{-1} \text{ sec}^{-1}$$

$$\text{At } 373.2^\circ\text{K} \quad \mu = (2.12 \times 10^{-4}) \exp (3.8) = 9.5 \times 10^{-3} \text{ g cm}^{-1} \text{ sec}^{-1}$$

Summary of results:

	0°C	100°C
Observed viscosity (μ)	1.787	0.2821
Predicted by Eq. 1.5-10	48.4	2.37
Predicted by Eq. 1.5-12	4.0	0.95

Both equations give poor predictions. This is not surprising inasmuch as the empirical formula in Eq. 1.5-11 et seq. do not hold for water, nor for most associated substances.

1.E Molecular Velocity and Mean Free Path

From Eq. 1.4-1 the mean speed of a molecule is

$$\bar{u} = \sqrt{\frac{8RT}{\pi M}} = \sqrt{\frac{8(8.314 \times 10^7)(273.2)}{\pi(32.00)}} = 4.25 \times 10^4 \text{ cm sec}^{-1}$$

From Eq. 1.4-3 the mean free path is:

$$\lambda = \frac{RT}{\sqrt{2} \pi d^2 p N} = \frac{(82.057)(273.2)}{\sqrt{2} \pi (3 \times 10^{-8})^2 (1)(6.023 \times 10^{23})} = 9.3 \times 10^{-6} \text{ cm}$$

Hence the mean free path is $(9.3 \times 10^{-6}) / (3 \times 10^{-8}) = 310$ molecular diameters under these conditions. In the liquid state, on the other hand, the corresponding ratio would be of the order of magnitude of, or even less than, unity.

1.F Comparison of the Uyehara-Watson Chart with Kinetic Theory

(a) Combination of Eqs. 1.4-11 and 18 gives:

$$\frac{\mu \sigma^2}{\sqrt{MT_c}} = 2.6693 \times 10^{-5} \frac{\sqrt{T_r}}{\Omega_{\mu}}$$

In addition, from Eq. 1.4-11 we know that $T_r = 0.77 RT/\epsilon$

The calculations for the kinetic theory plot are summarized below:

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Environmental Engineering Unit Operations and Unit Processes Association of Environmental Engineering Professors, John T. O'Connor, 1975

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McCabe, Julian Cleveland Smith, Peter Harriott, 1985 **Unit Operations of Chemical Engineering** Warren Lee
The Publishers' Trade List Annual , 1978

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