

Solutions Study Guide

6. What factor increases solubility of a solid?

7. Identify the

solubility

8. Identify the

9. 100 g

10. 100 g

11. 2.2 g

12. Identify the

13. 100 g

14. 100 g

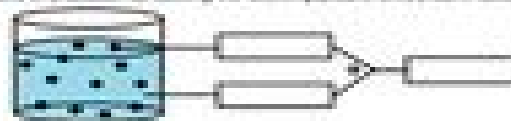
15. 100 g

16. 100 g

Name _____ Period _____ Date _____

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1. Identify the parts of a solution by labeling the following amount. (Solvent, Solute, Solution)



2. Circle which of the following are Solutions & Cross out the ones which are not solutions.

- ☐ Sugar
- ☐ Iron
- ☐ Salt Water
- ☐ Orange Juice
- ☐ Air
- ☐ Brass

3. Water plays an important role in creating solutions. It is known as the _____.

a. Universal Solvent

b. Universal Solvent

c. Why does water make such a good solvent?

d. Which type of solute would be soluble in water? (Remember "like dissolves like")

4. Order a word to match the word to its definition.

Solute	What is dissolved
Solvent	Substance that dissolves another substance in the solution
Solution	Substance that dissolves
Saturated	Substance that cannot dissolve any more
Unsaturated	Substance that can dissolve more
Supersaturated	Substance that can dissolve more than it should
Concentration	Amount of solute in a solution
Dilution	Process of making a solution less concentrated
Evaporation	Process of turning a liquid into a gas
Boiling	Process of turning a liquid into a gas at a higher temperature
Freezing	Process of turning a liquid into a solid
Melting	Process of turning a solid into a liquid

Solutions Study Guide Chemistry

David R. Klein



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