

Name: _____ Per: _____ Date: _____

Periodic Trends Worksheet

Use the periodic table and your knowledge of periodic trends to answer the following questions.

1. Which atom in each pair has the larger atomic radius (circle)?
a) Ca or Ni b) Ga or B c) O or C d) Cl or Br
2. What is the trend for atomic size from top to bottom in a group? From left to right in a period?
3. Why do atoms get smaller as you move left to right in a period?
4. Which element in each pair has a larger ionization energy (circle)?
a) Be or Ba b) Cu or Ra c) K or V d) W or Cr
5. Explain the relationship between the relative size of an ion to its neutral atom and the charge on the ions.
6. Which particle has the larger radius in each atom/ion pair (circle)?
a) Na, Na^+ b) S, S^{2-} c) I, I^- d) Al, Al^{3+}
7. What is ionization energy?
8. What is the periodic trend for ionization energy for a group? For a period?
9. What is the reason for the change in ionization energy as you go down a group?

Periodic Trends Worksheet Extension Questions

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



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Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **Resources in Education**, 1994

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Table of Contents Periodic Trends Worksheet Extension Questions

1. Understanding the eBook Periodic Trends Worksheet Extension Questions
 - The Rise of Digital Reading Periodic Trends Worksheet Extension Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Worksheet Extension Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Worksheet Extension Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Worksheet Extension Questions
 - Personalized Recommendations
 - Periodic Trends Worksheet Extension Questions User Reviews and Ratings
 - Periodic Trends Worksheet Extension Questions and Bestseller Lists
5. Accessing Periodic Trends Worksheet Extension Questions Free and Paid eBooks
 - Periodic Trends Worksheet Extension Questions Public Domain eBooks
 - Periodic Trends Worksheet Extension Questions eBook Subscription Services
 - Periodic Trends Worksheet Extension Questions Budget-Friendly Options

6. Navigating Periodic Trends Worksheet Extension Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Worksheet Extension Questions Compatibility with Devices
 - Periodic Trends Worksheet Extension Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Worksheet Extension Questions
 - Highlighting and Note-Taking Periodic Trends Worksheet Extension Questions
 - Interactive Elements Periodic Trends Worksheet Extension Questions
8. Staying Engaged with Periodic Trends Worksheet Extension Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Worksheet Extension Questions
9. Balancing eBooks and Physical Books Periodic Trends Worksheet Extension Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Worksheet Extension Questions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Worksheet Extension Questions
 - Setting Reading Goals Periodic Trends Worksheet Extension Questions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Worksheet Extension Questions
 - Fact-Checking eBook Content of Periodic Trends Worksheet Extension Questions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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