

Nuclear Reactions

Nuclear reactions are characterized by the following:

- They involve changes to the nucleus
 - Number of protons or neutrons in the nucleus change
- Energy changes are enormous
- Not concerned with valence electrons.
- Changes in temperature and pressure have no effect on the rate of the reaction
- Small particles are important.

Important subatomic particles

proton: ${}^1_1\text{p}$ ${}^1_1\text{H}$

neutron: ${}^1_0\text{n}$

Alpha Particle: ${}^4_2\alpha$ ${}^4_2\text{He}$ 2 protons
2 neutrons

beta particle: ${}^0_{-1}\beta$ ${}^0_{-1}\text{e}$ - high energy electrons emitted from the nucleus

positron: ${}^0_{+1}\beta$ ${}^0_{+1}\text{e}$ - positive electron: what gives a proton its positive charge

Gamma Ray: ${}^0_0\gamma$ - high energy photon

Definitions:

Isotopes - same number of protons but different mass number (mass number = # protons + # neutrons)

${}^{12}_6\text{C}$ ${}^{13}_6\text{C}$ ${}^{14}_6\text{C}$ are each an isotope of C (How many neutrons does each have?)

Nuclide (pronounced NOO clyde) - a nucleus with a specific number of protons and neutrons

Radioactive - Changes in matter resulting from SPONTANEOUS changes in the composition of the nucleus.

- Such changes are often accompanied by the creation of the subatomic particles above.
- Changes to the nucleus often result in the conversion of one element into another.

Radionuclide - nuclides that are radioactive.



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A Gutmann



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