

Chemistry

FORMULAE SHEET

$$n = \frac{m}{MM}$$

$$q = mc\Delta T$$

$$pK_a = -\log_{10}[K_a]$$

$$c = \frac{n}{V}$$

$$\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$$

$$A = eIc = \log_{10} \frac{I_0}{I}$$

$$PV = nRT$$

$$\text{pH} = -\log_{10}[\text{H}^+]$$

Avogadro constant, N_A	$6.022 \times 10^{23} \text{ mol}^{-1}$
Volume of 1 mole ideal gas: at 100 kPa and	
at 0°C (273.15 K)	22.71 L
at 25°C (298.15 K)	24.79 L
Gas constant	$8.314 \text{ J mol}^{-1} \text{ K}^{-1}$
Ionisation constant for water at 25°C (298.15 K), K_w	1.0×10^{-14}
Specific heat capacity of water	$4.18 \times 10^3 \text{ J kg}^{-1} \text{ K}^{-1}$

DATA SHEET

Solubility constants at 25°C

Compound	K_{sp}	Compound	K_{sp}
Barium carbonate	2.58×10^{-9}	Lead(II) bromide	6.60×10^{-6}
Barium hydroxide	2.55×10^{-4}	Lead(II) chloride	1.70×10^{-5}
Barium phosphate	1.3×10^{-29}	Lead(II) iodide	9.8×10^{-9}
Barium sulfate	1.08×10^{-10}	Lead(II) carbonate	7.40×10^{-14}
Calcium carbonate	3.36×10^{-9}	Lead(II) hydroxide	1.43×10^{-15}
Calcium hydroxide	5.02×10^{-6}	Lead(II) phosphate	8.0×10^{-43}
Calcium phosphate	2.07×10^{-29}	Lead(II) sulfate	2.53×10^{-8}
Calcium sulfate	4.93×10^{-5}	Magnesium carbonate	6.82×10^{-6}
Copper(II) carbonate	1.4×10^{-10}	Magnesium hydroxide	5.61×10^{-12}
Copper(II) hydroxide	2.2×10^{-20}	Magnesium phosphate	1.04×10^{-24}
Copper(II) phosphate	1.40×10^{-37}	Silver bromide	5.35×10^{-13}
Iron(II) carbonate	3.13×10^{-11}	Silver chloride	1.77×10^{-10}
Iron(II) hydroxide	4.87×10^{-17}	Silver carbonate	8.46×10^{-12}
Iron(III) hydroxide	2.79×10^{-39}	Silver hydroxide	2.0×10^{-8}
Iron(III) phosphate	9.91×10^{-16}	Silver iodide	8.52×10^{-17}
		Silver phosphate	8.89×10^{-17}
		Silver sulfate	1.20×10^{-5}

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R Barnett



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Table of Contents Ny State Chemistry Exam Formula Sheet

1. Understanding the eBook Ny State Chemistry Exam Formula Sheet
 - The Rise of Digital Reading Ny State Chemistry Exam Formula Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Ny State Chemistry Exam Formula Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ny State Chemistry Exam Formula Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ny State Chemistry Exam Formula Sheet
 - Personalized Recommendations
 - Ny State Chemistry Exam Formula Sheet User Reviews and Ratings
 - Ny State Chemistry Exam Formula Sheet and Bestseller Lists
5. Accessing Ny State Chemistry Exam Formula Sheet Free and Paid eBooks
 - Ny State Chemistry Exam Formula Sheet Public Domain eBooks
 - Ny State Chemistry Exam Formula Sheet eBook Subscription Services

- Ny State Chemistry Exam Formula Sheet Budget-Friendly Options
- 6. Navigating Ny State Chemistry Exam Formula Sheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Ny State Chemistry Exam Formula Sheet Compatibility with Devices
 - Ny State Chemistry Exam Formula Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ny State Chemistry Exam Formula Sheet
 - Highlighting and Note-Taking Ny State Chemistry Exam Formula Sheet
 - Interactive Elements Ny State Chemistry Exam Formula Sheet
- 8. Staying Engaged with Ny State Chemistry Exam Formula Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ny State Chemistry Exam Formula Sheet
- 9. Balancing eBooks and Physical Books Ny State Chemistry Exam Formula Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ny State Chemistry Exam Formula Sheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ny State Chemistry Exam Formula Sheet
 - Setting Reading Goals Ny State Chemistry Exam Formula Sheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ny State Chemistry Exam Formula Sheet
 - Fact-Checking eBook Content of Ny State Chemistry Exam Formula Sheet
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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