#### Metal oxides and hydroxides

Alumina (aluminum oxide)

- Amphoteric reacts with acids and bases
- With acids it acts as a base:
- Al<sub>2</sub>O<sub>3</sub> + 6HCl → 2AlCl<sub>3</sub> + 3H<sub>2</sub>O • With bases it acts as an acid:
- Al<sub>2</sub>O<sub>3</sub> + 6NaOH → 3H<sub>2</sub>O + 2Na<sub>3</sub>Al(OH)<sub>3</sub> (sodium aluminate)

Alumina is unreactive, uses of it include paint, sunscreens, and chromatography. Group 1 and 2 metal oxides

- Basic due to the small size of the oxide ion
- Group 1/2 metal oxides + water → metal hydroxides:

Group 1 eg: NaO + H<sub>2</sub>O → 2NaOH

Group 2 eg:

Market Name

-

arthur a

CaO + H<sub>2</sub>O → Ca(OH)<sub>2</sub>-

Used to resituation acidic soil

#### Other uses

Mg(OH)<sub>2</sub> – indigestion

Ca(OH)<sub>2</sub> – used to treat acidic effluent from factories:

(H,SO, + Ca(OH), → CaSO, + 2H,O)

Properties and uses of substances 1



Chairm Makes affice

# BTEC Applied Science Unit 5 - Chemistry

Key terms

Base - a compound that reacts with acids

dissociates (breaks up)

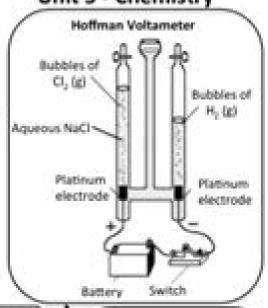
acid or a base.

Acid - a compound that releases H' ions when it

Alkali - a soluble base that releases OH in water

Amphoteric - a compound that can behave as an

Electrolysis - breaking down of a compound using



## Electrolysis

40,46

An electrical current must be able to flow, so electrolysis only works when compounds are either molten or dissolved

Used to make plastics and snop

#### Predicting products in aqueous electrolysis

In molten electrolysis there is only one +ve ion and one -ve ion, so predicting products is easy. In aqueous solution there are also H' and OH' ions, so we must predict which +ve and which -ve ion will discharge.

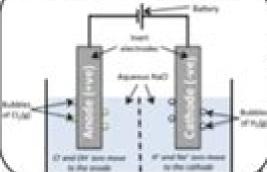
Factors affecting which ion is discharged:

- · Reactivity more reactive will discharge
- · Concentration more concentrated discharge
- Position in the electrochemical series (+ve ions only) lower in the series (higher/more positive standard electrode potential) more likely to discharge

### Electrolysis of aqueous NaCl (Brine)

CE Section Section In

**Electrolysis of molten NaCl** 



Specialised cells - H<sub>2</sub> and Cl<sub>2</sub> are explosive when mixed, and Cl<sub>2</sub> also reacts with NaCl so the products must be separate during the process

#### Membrane cell

Has a polymer membrane that only allows positive ions to pass through. Na' ions pass through but not Cl' ions.

#### Membrane cell

Has a diaphragm made of a porous mixture of asbestos and polymers. NaCl(aq) is added on one side and NaOH is made on the other. The level on the brine side is kept higher so it does not flow back and mix with the Cl<sub>2</sub>(g).

#### Uses of brine electrolysis products

NaOH – used in food processing and removing pollutants during paper manufacture

H<sub>2</sub> – used to make HCl and also as a fuel

Cl<sub>2</sub> – used as a disinfectant and to make plastics

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