

Acid Base Titration

1. Define the following terms.

Term	Definition
analyte	It is a substance whose concentration is to be determined during chemical reaction or procedure.
titrant	It is a solution of known concentration in titration, also known as titrator.
equivalence point	In titration, it is a point where the amount of added titrant solution is equal to titrand solution.
indicator	A substance which shows a colour change due to its specific pH range and gives end point.
end point	During titration, it is a point where indicator changes its colour.

2. Log onto <http://chemcollective.org/vlab/vlab.php> and start the Virtual Lab Aplet. Click on Help and watch the Video Walkthrough.

3. You will be doing a series of 3 different titrations.

Analyte	Analyte Volume	Titrant	Volume of Titrant Added Each Time	Total Volume of Titrant
0.10 M HCl	10.00 mL	0.10 M NaOH	0.50 mL	15.00 mL
0.10 M HC ₂ H ₃ O ₂ *	10.00 mL	0.10 M NaOH	0.50 mL	15.00 mL
0.10 M NH ₃ *	10.00 mL	0.10 HCl	0.50 mL	15.00 mL

You will need a 50 mL burette for your titrant and a 250 mL beaker for your analyte.

*You will be required to make 0.10 M concentrations of these analytes using the solutions in the cabinets, distilled water, a pipette, and a volumetric flask. Remember $c_1V_1=c_2V_2$

4. Create the table below for each titration. Your first entry should be a volume of 0 mL and **your final volume 15 mL**. Make sure you label each table, so you know which data goes with which titration.

Titration data for 0.10 M HCl and 0.10 M NaOH

Virtual Acid Base Ph Titration Packet Answers

Matt Kingsley



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LaMotte Chemical Products Co. Educational Products Division,1977? **Acid Base Titrations** Shah Purvesh,2015-08-04 In Acid Base Titrations the author discussed various acid base titration It gives some basic concept of acid base titration The analysis of acid base titration discussed in this book It also covers Titration of polyprotic acid and mixture of acids Titration of a Polybasic base with a Strong Acid pH transition range for an indicator titration of salts Differential Alkali titration are discussed in simplest type **General Theory of Acid-base Titration** Pat Richard Pondy,1949 **An On-line Acid-base Titration Applet in the Generic Tutorial System for the Sciences Project** Thomas Lee Gummo,2002 The purpose of this Master s Project was to develop an Acid Base Titration Simulator It was also to be a part of the California State University San Bernardino s GTSS Generic Tutorial System for the Sciences project The main benefit is that students will be able to conduct titration experiments over the Internet without being in the laboratory and without costly equipment or dangerous chemicals Instructors at the high school and college level can demonstrate the key principles of titration **Acid-Base Titration** Bird,1995-01-01 **Aqueous Acid-base Equilibria and Titrations** Robert De Levie,1999 This book will give students a thorough grounding in pH and associated equilibria material absolutely fundamental to the understanding of many aspects of chemistry It is in addition a fresh and modern approach to a topic all too often taught in an outmoded way This book uses new theoretical developments which have led to more generalized approaches to equilibrium problems these approaches are often simpler than the approximations which they replace Acid base problems are readily addressed in terms of the proton condition a convenient amalgam of the mass and charge constraints of the chemical system considered The graphical approach of Bjerrum Hagg and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria Based on these concentrations the proton condition can usually be simplified often leading directly to the value of the pH In the description of acid base titrations a general master equation is developed It provides a continuous and complete description of the entire titration curve which can then be used for computer based comparison with experimental data Graphical estimates of the steepness of titration curves are also developed from which the practicality of a given titration can be anticipated Activity effects are described in detail including their effect on titration curves The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements which are subject to activity corrections and balance equations and spectroscopic pH measurements which are not Finally an entire chapter is devoted to what the pH meter measures and to the experimental and theoretical uncertainties involved

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