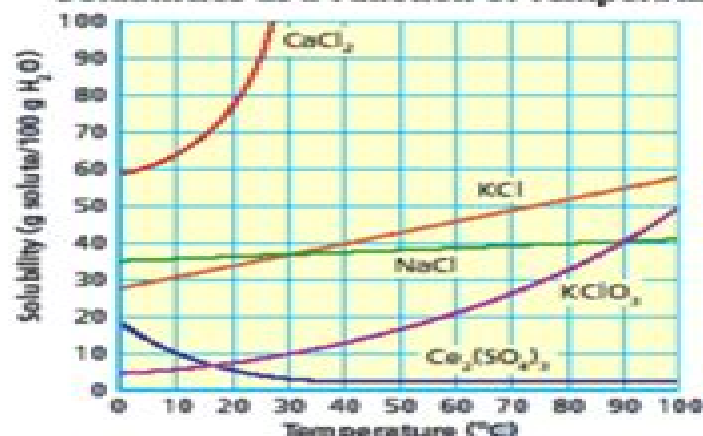


Name: \_\_\_\_\_

Date: \_\_\_\_\_ Period: \_\_\_\_\_

Solubility Curve Worksheet 2

Directions: Use the graph below to answer the following questions.

**Solubilities as a Function of Temperature****Types of Solutions**

- > Points that collectively make up the solubility curve (points **ON** the curve) represent \_\_\_\_\_ solutions.
- > Points that are **BELOW** the curve represent \_\_\_\_\_ solutions.
- > Points **ABOVE** the solubility curve represent \_\_\_\_\_ solutions, and the **difference** between the point above the curve and on the curve represents the amount of solute which will precipitate out.
- > Say if the resulting solutions would be saturated, supersaturated or unsaturated.
  - 60 g of KCl at 70°C \_\_\_\_\_
  - 10 g of KClO<sub>3</sub> at 60°C \_\_\_\_\_
  - 70 g of CaCl<sub>2</sub> at 20°C \_\_\_\_\_

1. a. What is the solubility of calcium chloride (CaCl<sub>2</sub>) at 5°C? \_\_\_\_\_  
 b. What is the solubility of calcium chloride at 25°C? \_\_\_\_\_  
 c. What is happening to the solubility and temperature for CaCl<sub>2</sub>? Solubility is \_\_\_\_\_ as temperature is \_\_\_\_\_.
2. a. What is the solubility of cerium sulfate - Ce<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> at 10°C? \_\_\_\_\_  
 b. What is the solubility of cerium sulfate at 50°C? \_\_\_\_\_  
 c. What is happening to the solubility and temperature for Ce<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>? Solubility is \_\_\_\_\_ as temperature is \_\_\_\_\_.
3. Which substance on the graph is **least** soluble at 10°C? \_\_\_\_\_
4. Which substance on the graph shows the **least** change in solubility from 0°C to 100°C? \_\_\_\_\_
5. Which solution is more concentrated:
  - a) At 10°C, a saturated solution of NaCl or a saturated solution of CaCl<sub>2</sub> (circle one)
  - b) At 55°C, a saturated solution of KClO<sub>3</sub> or a saturated solution of Ce<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> (circle one)
6. What is the mass of cerium sulfate that will dissolve in 50 g of water at 10 degrees Celsius?
7. What is the mass of potassium chloride that will dissolve in 250 g of water at 70 degrees Celsius?
8. At 90 degrees Celsius, 10 g of potassium chlorate is dissolved in 100 g of water. Is this solution saturated, unsaturated, or supersaturated? How do you know?
9. A saturated solution of potassium chlorate is dissolved in 100g of water. If the saturated solution is cooled from 90 degrees Celsius to 60 degrees Celsius, how many grams would crystallize out?

# Solubility Curve 1 Answer Key

**Bryan M. Ham,Aihui MaHam**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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