

Answer each item for 2 points each.

Define the following terms.

1. What is an isotope?
2. Define atomic mass in terms of subatomic particles.
3. What is the Aufbau Principle?
4. What is the dual nature of electrons?
5. What is STP?
6. Which isotope is the basis for the mole concept and relative atomic mass?

Explain the following chemical concepts.

7. Apply kinetic energy to explain how solids differ from liquids.
8. Is air matter? Explain.
9. Why do atoms form chemical bonds?
10. When Li and O bond, is it ionic, covalent, or metallic? Explain.
11. What is the relationship between wavelength and frequency?
12. What is the normal pH range for aqueous solutions?
13. Describe the equivalence point of an acid-base titration.
14. Provide an example of a diprotic acid.
15. What is the difference between a sulfide and sulfate ion?
16. What is the $[OH^-]$ of pure water at 25 $^{\circ}C$?
17. In a solution if the hydronium ion concentration is less than the hydroxide ion concentration, is the solution acidic, basic, or neutral?

Slo 2 Chemistry Post Test Answer Key

G Psacharopoulos



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