

# CHEMISTRY Terminology

The ABCs of chemistry

## A

**A:** Symbol for atomic mass number  
**absolute zero:** Lowest temperature; 0 Kelvin  
**Ae:** Symbol for the element actinium,  $Z = 89$   
**accuracy:** The closeness of measurement to true value  
**acetate anion:**  $C_2H_3O_2^-$ , charge -1  
**acetic acid:**  $C_2H_4O_2$   
**acetone:** Dimethyl ketone; organic solvent  
**acid-base indicator:** A chemical that changes color if the pH changes  
**acid-base titration:** A method for determining acid or base concentrations  
**acid definition:** See Arrhenius acid, Brønsted-Lowry acid, Lewis acid  
**acid ionization constant ( $K_a$ ):** Equilibrium constant for acid dissociation  
**acidic solution:** pH below 7  
**actinide:** Element with  $Z = 90$  or above; radioactive  
**activation energy**  
**( $E_a$ ):** A process's energy barrier



**adhesion:** Attraction of unlike molecules or materials  
**Ag:** Symbol for the element silver,  $Z = 47$   
**Al:** Symbol for the element aluminum,  $Z = 13$   
**alcohol:** Organic compound with -OH group; ROH  
**aldehyde:** Carbonyl group bonded to hydrogen and an organic group  
**alkali metal family:** Lithium, sodium, potassium, rubidium, cesium, and francium; column #1  
**alkaline earth metal family:** Beryllium, magnesium, calcium, strontium, barium, and radium; column #2  
**alkane:** Hydrocarbon; all C-C single bonds  
**alkene:** Hydrocarbon; 1 or more C=C double bonds  
**alkyne:** Hydrocarbon; 1 or more C≡C triple bonds  
**allotropes:** Two or more forms of an element  
**alloy:** Solution of 2 or more metals  
**alpha ( $\alpha$ ):** Greek letter that denotes radioactive particle and various scientific variables  
**alpha particle:** Helium nucleus; charge +2  
**Au:** Symbol for the element auricium,  $Z = 95$   
**amalgam:** Alloy of mercury with other metal  
**amine:** Organic base; RRRN; modified ammonia  
**amino acid:** Compound with organic acid and organic base property; forms proteins and peptides  
**ammonia:**  $NH_3$ , base  
**ammonium cation:**  $NH_4^+$ , charge +1  
**ammonium hydroxide:**  $NH_4OH$   
**amphoteric oxide:** Exhibits both acid and basic properties  
**amu:** Atomic mass unit;  $1/12$  mass of C-12  
**anion:** Ion with a negative charge  
**anode:** Electrode that supports oxidation  
**antibonding MO:** MO is less stable than separate AOs  
**AO:** Abbreviation for atomic orbital

**aq:** Aqueous; "dissolved in water"  
**Ar:** Symbol for the element argon,  $Z = 18$   
**aromatic:** Organic compound with a benzene ring  
**Arrhenius acid:** Produces hydronium ion in water solution  
**Arrhenius base:** Yields hydroxide ion in water solution  
**Arrhenius equation for rate constant ( $k$ ):**  

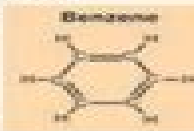
$$k = Ae^{-E_a/RT}$$
 $E_a$  = activation energy  
**arsenide anion:**  $As^{3-}$ , charge -3  
**As:** Symbol for the element arsenic,  $Z = 33$   
**At:** Symbol for the element astatine,  $Z = 85$   
**atm:** Symbol for pressure in "atmospheres"  
**atom:** Fundamental unit of all matter  
**atomic mass number ( $A$ ):** Total number of protons and neutrons  
**atomic number ( $Z$ ):** Number of protons in the nucleus  
**atomic orbital:** Wave motion of electrons in atoms  
**atomic radius:** Empirical measure of atom size  
**atomic weight:** Weighted average of natural isotopes of an element  
**Au:** Symbol for the element gold,  $Z = 79$   
**Aufbau principle:** Guides the filling of electronic subshells of the elements  
**Avogadro's law:** Volume is proportional to gas moles at fixed pressure and temperature  
**Avogadro's number:**  $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$

**Bk:** Symbol for the element berkelium,  $Z = 97$   
**bleach:** Chemical that is a strong oxidizing agent  
**boiling:** Liquid  $\rightarrow$  gas at the boiling point  
**boiling point elevation:** A solution has a higher boiling point than a pure solvent  
**boiling point ( $T_b$ ):** Liquid-gas equilibrium,  $P = 1 \text{ atm}$   
**boiling point of water:** 100°C  
**bond energy:** Energy held by a chemical bond  
**bond length:** Distance between 2 atoms in a bond  
**bond order for Lewis model:** # of bonds divided by # of bonded atoms  
**bond order for MO treatment:** # of filled bonding MOs minus # of filled antibonding MOs  
**bonding electrons:** Form chemical bonds in a compound  
**bonding MO:** MO is more stable than AOs  
**Boyle's law:**  $PV$  constant for gas, fixed  $T$  and  $n$   
**Br:** Symbol for the element bromine,  $Z = 35$   
**bromide anion:**  $Br^-$ , charge -1  
**Brønsted-Lowry acid:** Proton donor  
**Brønsted-Lowry base:** Proton acceptor  
**buffer:** Composed of weak acid and weak base; serves to keep pH constant



## B

**B:** Symbol for the element boron,  $Z = 5$   
**Ba:** Symbol for the element barium,  $Z = 56$   
**balanced equation:** Each side of the equation has the same number of atoms of each element and charge  
**balanced redox:** Electron loss = electron gain  
**barium cation:**  $Ba^{2+}$ , charge +2  
**base definition:** See Arrhenius base, Brønsted-Lowry base, Lewis base  
**base ionization constant ( $K_b$ ):** Equilibrium constant for base dissociation  
**basic solution:** pH over 7  
**battery:** Source of electrical power; galvanic cell  
**Be:** Symbol for the element beryllium,  $Z = 4$   
**benzene:**  $C_6H_6$ ; organic solvent; aromatic ring of 6 carbon atoms  
**beryllium cation:**  $Be^{2+}$ , charge +2  
**beta ( $\beta$ ):** Greek letter that denotes radioactive particle and a number of scientific variables  
**beta particle:** Energetic electron from the nucleus; charge -1  
**Bh:** Symbol for the element bohrium,  $Z = 107$   
**Bi:** Symbol for the element bismuth,  $Z = 83$   
**bicarbonate anion:**  $HCO_3^-$ , charge -1  
**bimolecular oxygen:**  $O_2$   
**binary compound:** Formed from 2 elements  
**bisulfate anion:**  $HSO_4^-$ , charge -1



**b:** Symbol for  $\text{centi-}$ , the SI prefix for  $1/100$  or  $10^{-2}$   
**c:** Symbol for the speed of light  
**°C:** Celsius temperature scale  
**C:** Symbol for the element carbon,  $Z = 6$   
 **$C_2H_4O_2$ :** Acetic acid  
**Ca:** Symbol for the element calcium,  $Z = 20$   
**calcium cation:**  $Ca^{2+}$ , charge +2  
**calcium fluoride:**  $CaF_2$   
**calcium hydroxide:**  $Ca(OH)_2$   
**calorie:** English unit of heat energy  
**calorimetry:** Study of the release or absorption of heat  
**carbide anion:**  $C^{4-}$ , charge -4  
**carbohydrate:** Organic compound; carbon bonded to several -OH groups (sugar and starch)  
**carbon dioxide:**  $CO_2$   
**carbon monoxide:**  $CO$   
**carbonate anion:**  $CO_3^{2-}$ , charge -2  
**carbonic acid:**  $H_2CO_3$   
**carbonyl group:**  $CO$ , ligand complex  
**carboxyl group:**  $-CO$  bonded to 2 groups  
**carboxylic acid:** Organic acid;  $R-COOH$   
**catalyst:** Accelerates a reaction but is not consumed in the reaction; lowers the activation energy  
**cathode:** Electrode that supports reduction  
**cation:** Ion with a positive charge  
**Cd:** Symbol for the element cadmium,  $Z = 48$



# Slo Test Study Guide For Chemistry

**R Sanford**



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$v_1$  is the supply voltage per phase when the full voltage  $v_1$  per phase is applied to the direct switching the starting current drawn from the supply is given by the equation shown below with auto transformer starting if tapping of the transformer ratio  $x$  is used then the voltage per phase across the motor is  $xv_1$

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