

Electron Configuration of Main Group Cations

Electron configuration

Na $1s^2 2s^2 2p^6 3s^1$

Orbital diagram



Focus on the valence shell

The charge is +1, therefore, one electron is lost:



The electron must be from the outermost valence orbital – $3s^1$



Na $1s^2 2s^2 2p^6 3s^1$

Na⁺ $1s^2 2s^2 2p^6$

↓ Same as

Ne $1s^2 2s^2 2p^6$



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Cations of main group metals attain the electron configuration of the noble gas that precedes them in the periodic table. They are **isoelectronic** with the previous noble gas.

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**United States. Energy Research and
Development Administration.
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Nuclear Science Abstracts ,1971 **Chemistry** Eugene LeMay, Jr.,Herbert Beall,Karen M. Robblee,Douglas C. Brower,Prentice-Hall Staff,2002-02 **ERDA Energy Research Abstracts** United States. Energy Research and Development Administration,1976 Scientific and Technical Aerospace Reports ,1988 **Congressional Record** United States. Congress,1978 **Energy Research Abstracts** ,1991-10 *Fusion Energy Update* ,1978 *Government-wide Index to Federal Research & Development Reports* ,1967 *ERDA Energy Research Abstracts* United States. Energy Research and Development Administration. Technical Information Center,1976 **Plasma Physics and Magnetohydrodynamics** Defense Documentation Center (U.S.),1962 **Physics Briefs** ,1994 *U.S. Government Research Reports* ,1963 *Using high energy density plasmas for nuclear experiments relevant to nuclear astrophysics* Maria Gatu Johnson,Michael Wiescher,Mark Paris,Alex Zylstra,Gerald Hale,2023-04-14 *U.S. Government Research & Development Reports* ,1967 ERDA Energy Research Abstracts United States. Energy Research and Development Administration,1976-11 **Government Reports Announcements & Index** ,1993 **Encyclopaedia Britannica** Harry S. Ashmore,1962 **INIS Atomindex** ,1987 1992 Fall Meeting ,1992 Government Reports Announcements ,1973

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