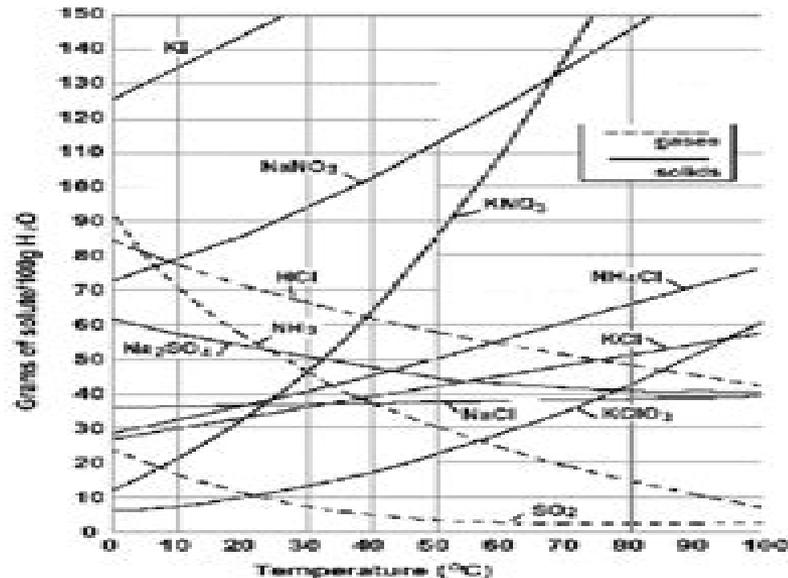


Lesson 4.1 Solubility Curves



Directions: Use the graph to answer the questions below. If the question requires a calculation, **SHOW ALL WORK TO RECEIVE CREDIT FOR THE ANSWER!**

- What is the solubility of potassium chloride at 10°C? 30g/100g H₂O
- What is the solubility of potassium chloride at 80°C? ~52g/100g H₂O
- Is potassium chloride a solid or a gas? solid What happens to the solubility of solids as temperature increases?
When temp increases solubility increases
- What is the solubility of NH₃ at 10°C?
70g/100g H₂O
- What is the solubility of NH₃ at 50°C?
30g/100g H₂O
- Is ammonia a solid or gas? gas What happens to the solubility of gases as temperature increases?
When temp increases solubility decreases
- What is the mass of potassium chlorate that will dissolve in 50 g of water at 20°C? 5g/50g H₂O
Solubility(20degree C) = 10g/100g H₂O

$$\frac{x \text{ g}}{50 \text{ g H}_2\text{O}} = \frac{10 \text{ g}}{100 \text{ g H}_2\text{O}} = 5\text{g}/50\text{g H}_2\text{O} \rightarrow 5\text{g}/50\text{g H}_2\text{O}$$
- What is the mass of sodium sulfate that will dissolve in 250 g of water at 30°C? 130g/250g H₂O

$$\frac{250}{100} = 2.5$$

$$51\text{g}/100\text{g H}_2\text{O} \rightarrow 51 \times 2.5 = 127.5 = 130\text{g}$$
- At 90°C, 10 g of potassium chlorate is dissolved in 100 g of water. Is this solution saturated, unsaturated, or supersaturated? undersaturated
 @ 90 degree c = 52g/100g H₂O -> 10g < 52g
- A saturated solution of potassium chlorate is dissolved in 100 g of water. If the saturated solution is cooled from 90°C to 60°C, how many grams of precipitate will be formed?
 @ 90 degree c = 52g/100g H₂O
 @ 60 degree c = 28g/100g H₂O 52g - 28g = 24g

Using Solubility Curves Answers

**Barron's Educational Series, Albert
Tarendash**



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Regents Exams and Answers: Chemistry--Physical Setting Revised Edition Barron's Educational Series, Albert Tarendash, 2021-01-05 Barron's Regents Exams and Answers Chemistry provides essential practice for students taking the Chemistry Regents including actual recently administered exams and thorough answer explanations for all questions This book features Eight actual administered Regents Chemistry exams so students can get familiar with the test Thorough explanations for all answers Self analysis charts to help identify strengths and weaknesses Test taking techniques and strategies A detailed outline of all major topics tested on this exam A glossary of important terms to know for test day

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