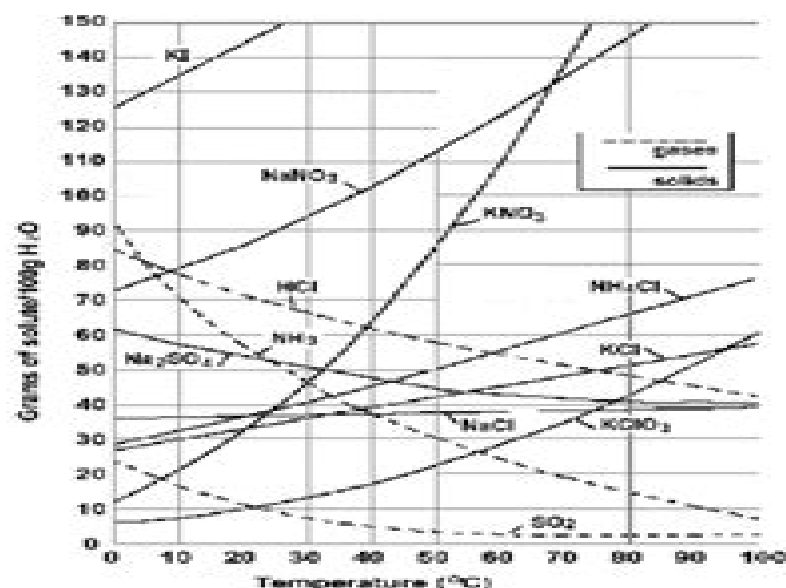


# Lesson 4.1 Solubility Curves



**Directions:** Use the graph to answer the questions below. If the question requires a calculation, **SHOW ALL WORK TO RECEIVE CREDIT FOR THE ANSWER!**

- What is the solubility of potassium chloride at 10°C? 30g/100g H<sub>2</sub>O
- What is the solubility of potassium chloride at 80°C? ~52g/100g H<sub>2</sub>O
- Is potassium chloride a solid or a gas? solid What happens to the solubility of solids as temperature increases?  
When temp increases solubility increases
- What is the solubility of NH<sub>3</sub> at 10°C? 70g/100g H<sub>2</sub>O
- What is the solubility of NH<sub>3</sub> at 50°C? 30g/100g H<sub>2</sub>O
- Is ammonia a solid or gas? gas What happens to the solubility of gases as temperature increases?  
When temp increases solubility decreases
- What is the mass of potassium chlorate that will dissolve in 50 g of water at 20°C? 5g/50g H<sub>2</sub>O  
Solubility(20degree C) = 10g/100g H<sub>2</sub>O  

$$\frac{x \text{ g}}{50 \text{ g H}_2\text{O}} = \frac{10 \text{ g}}{100 \text{ g H}_2\text{O}} = 5\text{g}/50\text{g H}_2\text{O}$$
- What is the mass of sodium sulfate that will dissolve in 250 g of water at 30°C? 130g/250g H<sub>2</sub>O  

$$\frac{250}{100} = 2.5$$

$$51\text{g}/100\text{g H}_2\text{O} \rightarrow 51 \times 2.5 = 127.5 = 130\text{g}$$
- At 90°C, 10 g of potassium chlorate is dissolved in 100 g of water. Is this solution saturated, unsaturated, or supersaturated? undersaturated  
@ 90 degree c = 52g/100g H<sub>2</sub>O -> 10g < 52g
- A saturated solution of potassium chlorate is dissolved in 100 g of water. If the saturated solution is cooled from 90°C to 60°C, how many grams of precipitate will be formed?  
@ 90 degree c = 52g/100g H<sub>2</sub>O  
@ 60 degree c = 28g/100g H<sub>2</sub>O     52g - 28g = 24g

# Using Solubility Curves Answers

**Barron's Educational Series, Albert  
Tarendash**



## Using Solubility Curves Answers:

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## **Table of Contents Using Solubility Curves Answers**

1. Understanding the eBook Using Solubility Curves Answers
  - The Rise of Digital Reading Using Solubility Curves Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Using Solubility Curves Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Using Solubility Curves Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Using Solubility Curves Answers
  - Personalized Recommendations
  - Using Solubility Curves Answers User Reviews and Ratings
  - Using Solubility Curves Answers and Bestseller Lists

5. Accessing Using Solubility Curves Answers Free and Paid eBooks
  - Using Solubility Curves Answers Public Domain eBooks
  - Using Solubility Curves Answers eBook Subscription Services
  - Using Solubility Curves Answers Budget-Friendly Options
6. Navigating Using Solubility Curves Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Using Solubility Curves Answers Compatibility with Devices
  - Using Solubility Curves Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Using Solubility Curves Answers
  - Highlighting and Note-Taking Using Solubility Curves Answers
  - Interactive Elements Using Solubility Curves Answers
8. Staying Engaged with Using Solubility Curves Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Using Solubility Curves Answers
9. Balancing eBooks and Physical Books Using Solubility Curves Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Using Solubility Curves Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Using Solubility Curves Answers
  - Setting Reading Goals Using Solubility Curves Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Using Solubility Curves Answers
  - Fact-Checking eBook Content of Using Solubility Curves Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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