

Stoichiometry: Assessment Review

Name _____

Skill – Student can use a balanced equation to determine mole to mole ratios

- 1) Given this equation: $N_2 + 3 H_2 \rightarrow 2 NH_3$, write the following molar ratios:
 - a. $N_2 : H_2$
 - b. $N_2 : NH_3$
 - c. $H_2 : NH_3$
- 2) Given the following equation: $8 H_2 + S_8 \rightarrow 8 H_2S$, write the following molar ratios:
 - a. $H_2 : H_2S$
 - b. $H_2 : S_8$
 - c. $H_2S : S_8$
- 3) Answer the following questions for this equation: $2 H_2 + O_2 \rightarrow 2 H_2O$
 - a. What is the $H_2 : H_2O$ molar ratio?
 - b. With 20 moles of H_2 on hand & plenty of O_2 , how many moles of H_2O could you make?
 - c. What is the $O_2 : H_2O$ molar ratio?
 - d. With 20 moles of O_2 and enough H_2 , how many moles of H_2O could you make?
- 4) Answer the following questions for this equation: $N_2 + 3 H_2 \rightarrow 2 NH_3$
 - a. If you used 1 mole of N_2 , how many moles of NH_3 could be produced?
 - b. If 10 moles of NH_3 were produced, how many moles of N_2 would be required?
 - c. If 3.00 moles of H_2 were used, how many moles of NH_3 would be made?
 - d. If 0.000 moles of NH_3 were produced, how many moles of H_2 are required?

Skill – Student can use a balanced equation, given the grams of reactant, to calculate the grams of another substance

- 5) Magnesium reacts with hydrochloric acid according to the following equation:



- a) If 72.92 grams of hydrochloric acid react with excess magnesium, how many moles of hydrogen gas will be produced?
- b) Aluminum reacts with hydrobromic acid to produce aluminum bromide and hydrogen gas. Write a balanced equation for the reaction and calculate the number of moles of hydrobromic acid required to react with 29.47 grams of aluminum.

Review Stoichiometry Answer Key

Steven S Zumdahl



Review Stoichiometry Answer Key:

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