# Module 2: Alcohols, Halogenoalkanes and Analysis

(a) Explain, in terms of hydrogen bonding, the water solubility and the relatively low volatility of alcohols;

Because Alcohols have -OH groups attached to them, they can form hydrogen bonds,

Ethanol has Hydrogen bonds

Hydrogen bonds are stronger

More energy required to overcome Hydrogen bonds

Less energy required to overcome Van der Waals forces

Dissolve in water because they have -OH groups attached to them, so can form hydrogen bonds with water

High Volatility	>	Turns into Vapour easily	 Low boiling Point
Low Volatility	************	Difficult to turn Into vapour	 High boiling Point

(b) Describe the industrial production of ethanol by:

(i) fermentation from sugars, i.e. from glucose,

(iii) the reaction of ethene with steam in the presence of an acid catalyst:

Hydration of Ethene	Fermentation of Glucose
	Conditions:
300°C	30-35°C
H <sub>1</sub> PO <sub>4</sub>	Yeast
60atm pressure	Anaerobic/Lack of oxygen
C <sub>2</sub> H <sub>4</sub> + H <sub>2</sub> O> C <sub>2</sub> H <sub>4</sub> O	C.H. O. > 2C.H.O + 2CO.

(c) Outline, for alcohols:

(i) The use of ethanol in alcoholic drinks and as a solvent in the form of methylated spirits,

As solvent – e.g. methylated spirits
As petrol additive - To improve actane number e.g. Methanal
As feedstack – Used to make other organic compounds e.g. alcohol to alkenes

(d) Classify alcohols into primary, secondary and tertiary alcohols;

Primary Alcohol – The –OH group is attached to a carbon with 1 or no methyl group attached.

Secondary Alcohol - The –OH group is attached to a carbon with 2 methyl groups attached.

Tertiary Alcohol - The –OH group is attached to a carbon with 3 methyl groups attached.

(e) Describe the combustion of alcohols;

Alcohols are combusted by oxidation.

An Alcohol is combusted completely if oxidised in an excess of Oxygen

It forms Carbon dioxide and water if this occurs.

An Alcohol is combusted incompletely if oxidised when there is a shortage of oxygen

It forms Carbon monoxide and water.

(g) Describe the oxidation of alcohols using  $Cr_2O_7^{2-}/H^*$  (i.e.  $K_2Cr_2O_7/H_2SO_4$ ), including:

Oxidation – Either Adding O<sub>2</sub> or removing H<sub>2</sub> – Remove H from OH group and Remove H from the Carbon with the OH group With acidified Potassium Dichromate Colour change from Orange to green  $(K_2Cr_2O_7 \rightarrow Cr_3^+)$ 

(i) The oxidation of primary alcohols to form: aldehydes carboxylic acids; the control of the oxidation product using different reaction conditions,

Aldehyde – Formed from distillation of primary alcohol

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**Rob Ritchie, Emma Poole** 

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