WW: Periodic Trends Worksheet Name Score: /56
Use the periodic table and your knowledge of periodic trends to answer the following questions.

Atomic Radius

f) Be or Ba g) Si or S h) Fe or Au

- Why do atoms get smaller as you move left to right in a period? Larger as you move down a group?
- Arrange the following in order of increasing atomic radius atomic radius.
 Mg, Na, P, Si and Ar

lonic Radius

 Explain the relationship between the relative size of an ion to its neutral atom and the charge on the ions.

5. Which particle has the larger radius in each atom/ion pair?
a) Na, Na+
b) S, S²⁻
c) I, I⁻
d) AI, AI³⁺
e. CI, Cl-

Ionization Energy

6. Which element in each pair has the larger ionization energy?

g) Ca or Fr h) W or Se

a) Na or O b) Be or Ba c) Ar or F d) Cu or Ra e) I or Ne

Electronegativity

f) K or V

8. Which element in each pair has a higher electronegativity value?
a) Cl, F
b) C, N
c) Mg, Ne
d) As, Ca
e) Fr, Cs

Metallic Character

9. Arrange the elements from least to most metallic: Sodium, barium, cesium, magnesium, calcium

Electron Affinity

10. Write the chemical equation for the addition of an electron to a chlorine atom.

11. Arrange the following atoms in order of increasing electron affinity: Br, Sb, I, Te, Cl

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