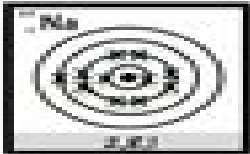

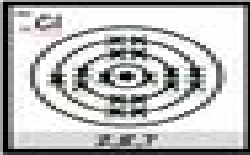
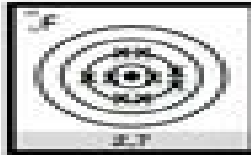



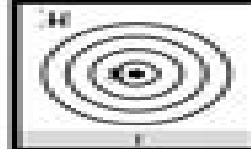


Formation of Ions

1. Use the element symbols to draw in the electronic configuration for each atom.
2. Determine how many electrons the atom needs to gain/lose.
3. Fill in the new electronic configuration for the ion.
4. Label the charge on the ion. The first one has been done for you.

	<p>Sodium</p> <p>Sodium needs to lose one electron.</p>		<p>Potassium</p>
	<p>Chlorine</p>		<p>Fluorine</p>
	<p>Magnesium</p>		<p>Aluminium</p>
	<p>Oxygen</p>		<p>Hydrogen</p>

Teaching Transparency Workseet Ionic Bonds

JA Banks



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