

**Model Question Paper-1 with effect from 2018-19  
(CBCS Scheme)**

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18CHE12/22

**First/Second Semester B.E.Degree Examination  
Engineering Chemistry**

(Common to all Branches)

Time : 3 Hrs

Max.Marks:100

**Note: Answer any FIVE full questions, choosing one full question from each module**

**Module-1**

1. a. Define the terms: (i) Free energy, (ii) Entropy & (iii) Cell potential (6 Marks)
- b. What are concentration cells? The cell potential of Cu concentration cell  $\text{Cu}/\text{CuSO}_4 (0.005\text{M}) // \text{CuSO}_4 (X) / \text{Cu}$  is 0.0295 V at 25° C. Write the cell reaction and calculate the value of X. (7 Marks)
- c. What are batteries? Explain the construction and working of Li-ion battery, mention its applications. (7 Marks)

**OR**

2. a. Write short notes on primary, secondary and reserve batteries. (6 Marks)
- b. Explain the construction and working of Ni-MH battery. Mention its applications. (7 Marks)
- c. For the cell,  $\text{Fe}/\text{Fe}^{2+} (0.01\text{M}) // \text{Ag}^+ (0.1\text{M}) / \text{Ag}$  write the cell reaction and calculate the emf of the cell at 298K, if standard electrode potentials of Fe and Ag electrodes are -0.44V and 0.8 V respectively. (7 Marks)

**Module-2**

3. a. What is metallic corrosion? Describe the electrochemical theory of corrosion taking iron as an example. (7 Marks)
- b. Explain: (i) Differential metal corrosion & (ii) Water-line corrosion (6 Marks)
- c. What is electroplating? Explain the electroplating of chromium (7 Marks)

**OR**

4. a. What is meant by metal finishing? Mention (any five) technological importance of metal finishing. (6 Marks)
- b. What is electrolessplating? Explain the electrolessplating of copper. (8 Marks)
- c. Explain the factors affecting the rate of corrosion (i) Nature of corrosion product, (ii) Ratio of anodic to cathodic areas & (iii) pH (6 Marks)

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**Karin Nielsen-Saines**



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## **Table of Contents Vtu 3rd Sem Previous Year Question Paper**

1. Understanding the eBook Vtu 3rd Sem Previous Year Question Paper
  - The Rise of Digital Reading Vtu 3rd Sem Previous Year Question Paper
  - Advantages of eBooks Over Traditional Books
2. Identifying Vtu 3rd Sem Previous Year Question Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Vtu 3rd Sem Previous Year Question Paper
  - User-Friendly Interface
4. Exploring eBook Recommendations from Vtu 3rd Sem Previous Year Question Paper
  - Personalized Recommendations
  - Vtu 3rd Sem Previous Year Question Paper User Reviews and Ratings
  - Vtu 3rd Sem Previous Year Question Paper and Bestseller Lists
5. Accessing Vtu 3rd Sem Previous Year Question Paper Free and Paid eBooks

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  - Vtu 3rd Sem Previous Year Question Paper eBook Subscription Services
  - Vtu 3rd Sem Previous Year Question Paper Budget-Friendly Options
6. Navigating Vtu 3rd Sem Previous Year Question Paper eBook Formats
- ePub, PDF, MOBI, and More
  - Vtu 3rd Sem Previous Year Question Paper Compatibility with Devices
  - Vtu 3rd Sem Previous Year Question Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Vtu 3rd Sem Previous Year Question Paper
  - Highlighting and Note-Taking Vtu 3rd Sem Previous Year Question Paper
  - Interactive Elements Vtu 3rd Sem Previous Year Question Paper
8. Staying Engaged with Vtu 3rd Sem Previous Year Question Paper
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Vtu 3rd Sem Previous Year Question Paper
9. Balancing eBooks and Physical Books Vtu 3rd Sem Previous Year Question Paper
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Vtu 3rd Sem Previous Year Question Paper
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Vtu 3rd Sem Previous Year Question Paper
- Setting Reading Goals Vtu 3rd Sem Previous Year Question Paper
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Vtu 3rd Sem Previous Year Question Paper
- Fact-Checking eBook Content of Vtu 3rd Sem Previous Year Question Paper
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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