

LAB 4: PRECIPITATION REACTIONS

Complete a series of 10 double-replacement reactions between ionic compounds in aqueous solution. Observe the formation of precipitates and write the balanced chemical equation for each reaction.

Goals

1. Observe the formation of precipitates in double-replacement reactions.
2. Identify the compounds that form precipitates and write the balanced chemical equation for each reaction.

Materials and Equipment

Reagents

- 0.1 M calcium chloride, CaCl_2
- 0.1 M sodium carbonate, Na_2CO_3
- 0.1 M sodium nitrate, NaNO_3
- 0.1 M potassium phosphate, K_3PO_4
- 0.1 M potassium hydroxide, KOH
- 0.1 M calcium acetate, $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$
- 0.1 M sodium sulfate, Na_2SO_4
- 0.1 M silver nitrate, AgNO_3

Equipment: Beakers, test tubes

Other Supplies

Method/Procedure

Use the double-replacement reaction as a general type of reaction. Use the following as the reaction between sodium nitrate and sodium sulfate:



Use the double-replacement reaction as a general type of reaction. Use the following as the reaction between sodium nitrate and sodium sulfate:

Write the balanced chemical equation for each reaction.

Precipitation Reactions Lab Quiz Answer Key

Judd E. Hollander



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