# Stoichiometry of a Precipitation Reaction – Lab Report Assistantw

#### Exercise 1: Stoichiometry and a Precipitation Reaction

Data Table 1. Stoichiometry Values

Initial: CaCl <sub>2</sub> •2H <sub>2</sub> O (g)	1.000
Initial: CaCl <sub>2</sub> •2H <sub>2</sub> O (moles)	.00080mor
Initial: CaCl <sub>2</sub> (moles)	100 BRUDO
Initial: Na <sub>2</sub> CO <sub>3</sub> (moles)	.00680a
Initial: Na <sub>2</sub> CO <sub>3</sub> (g)	.7219
Theoretical: CaCO <sub>3</sub> (g)	0.681
Mass of Filter paper (g)	1.009
Mass of Filter Paper + CaCO <sub>3</sub> (g)	1.420
Actual: CaCO <sub>3</sub> (g)	. 424
% Yield:	(01.609/0

#### Questions:

- A) A perfect percent yield would be 100%. Based on your results, describe your degree of accuracy and suggest possible sources of error.
- B) What impact would adding twice as much Na<sub>2</sub>CO<sub>3</sub> than required for stoichiometric quantities have on the quantity of product produced?
- C) Determine the quantity (g) of pure CaCl<sub>2</sub> in 7.5 g of CaCl<sub>2</sub>•9H<sub>2</sub>O.
- D) Determine the quantity (g) of pure MgSO<sub>4</sub> in 2.4 g of MgSO<sub>4</sub>•7H<sub>2</sub>O.
- E) Conservation of mass was discussed in the background. Describe how conservation of mass (actual, not theoretical) could be checked in the experiment performed.

## **Precipitation Reactions Lab Report Answers**

**Daniel F McAuley** 

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