Answers for Predicting Products of Chemical Reactions

For all combustion reactions of hydrocarbons or hydrogen-carbon-oxygen molecules, the products will always be CO: and H:O (assuming a complete reaction).

a. Combustion:	C_6H_{12}	+	901	-	6CO ₃	*	6H;O
b. Combustion:	2C ₄ H ₆	+	1102	\rightarrow	8CO2	+	6H:O
c. Combustion:	$C_0H_{10}O_3$	94	707		6CO ₂		5H ₂ O

1. Synthesis: Mg + I₂ → MgI:

Note that Mg can only form Mg2 and I can only form I', so the product will must have a 1:2 cation anion ratio.

2. Double displacement: CuCl₂ + H₂S → CuS + 2HCl Note that the product is **not** H₂Cl₂. It is important to recognize that CuCl₂ is made of three ions, Cu¹⁺ and two Cl².

Double displacement: NaOH + HCIO₄ → NaCIO₄ + H₂O
 In this question, you must recognize that perchlorate, ClO₄, and hydroxide, OH¹, are polyatomic ions and will not break apart. Also, this is an acid-base reaction, so the products should be salt and water.

Decomposition: ZnCO₂ + heat → ZnO + CO₂.
 When reactions have heat as a reactant, it is very likely that they will involve decompositions. Carbonate compounds usually decompose to CO₂ and a metal oxide.

5. Single replacement: 2HCl + Zn → ZnCl₂ + H₂
Note that one reactant is in its elemental form; if a single replacement reaction is going to occur, the species NOT in its elemental form in the reactants (H¹ in this case), will end up in its elemental form in the products (H₂) and the species that IS in its elemental form (Zn) will end up ionized (Zn²). Note that zinc can only form a Zn² ion, so it will have two chloride ions. Note also that hydrogen in its elemental form is H₂, not H.

6. <u>Single replacement</u> 2Na + MgCl₂ → 2NaCl + Mg Again notice that one species is in its elemental form (Na). The magnesium in MgCl₂ is an ion (Mg²⁺), but is transformed into its elemental state (Mg), while the Na is converted into an ion (Na⁺, sodium only forms a +1 ion).

Double replacement CaCl: + K:CO → CaCO: + 2KCl
 Recognize that carbonate is a polyatomic ion (CO₁²) and that the cations are already stable ions (Ca²² and K²).

8. <u>Synthesis</u> 2K + CI: → 2KCI

Note that both materials are elemental species, so the only result could be a synthesis reaction. In the product, K^{*} and CF are formed. Note that we would NOT write the product as K₂Cl₂. Ionic compounds are written in their simplest forms.

Double replacement 3BaCl: + 2K.PO₁ → Bas(PO₁): + 6KCl
 Note that phosphate (PO₁) is a polyatomic ion and will not break apart. Since barium is a +2 ion, the barium phosphate will have a 3:2 ratio of Ba:PO₁ in order to balance the charge.

10. <u>Double replacement</u> H:SO_a + 2KOH → 2H:O + K:SO_a Notice that sulfate (SO_a²) is a polyatomic ion and that potassium exists as a +1 ion (K²). Also, this is an acid-base reaction, so the products should be salt and water.

11. Decomposition Al₂(CO₂): + beat → Al₂O₂ + 3CO₂
Another decomposition that will generate CO₂ and a metal oxide (note that Al forms a +3 ion, Al²⁻, and monatomic oxygen will have a -2 charge, O²⁻).

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