

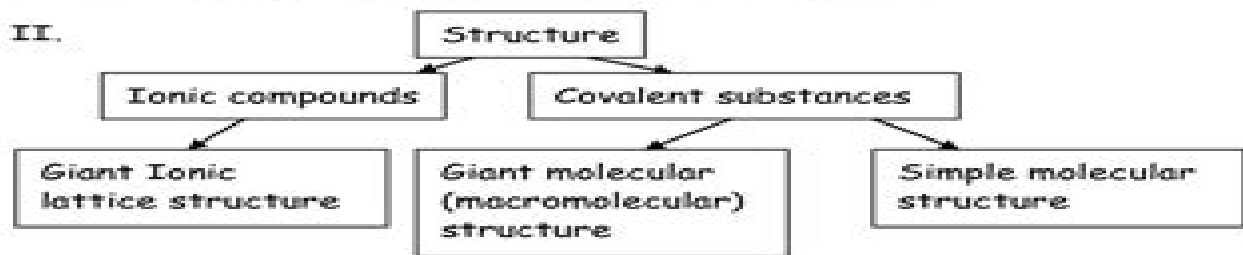
Answers to Worksheet 5 - Chemical Bonding
Science(Chemistry) Theory Workbook Pg 36 to 45

Check It (37 to 39)

1. C	2. B	3. D	4. A	5. B	6. A	7. D	8. B	9. B
------	------	------	------	------	------	------	------	------

- 10a) calcium oxide and potassium chloride
 b) carbon dioxide and methane
 c) calcium oxide or potassium chloride (any one, an ionic one)
 d) carbon dioxide or methane (any one, a covalent one)
 e) helium (a noble gas , is monatomic, as single atoms)
- 11a) C :(2,4) : O :(2,6) : Na :(2,8,1) : Si :(2,8,4) : Cl :(2,8,7)
 b)(i) No. of neutrons = Nucleon no – no of protons = 9 – 4 = 5
 (ii) No. of neutrons = 28 – 14 = 14
 c) Beryllium and sodium (are metals, but, diamond does not conduct)
 d)(i) NaCl, sodium chloride
 (ii) Na₂O, sodium oxide
 (iii) CO₂, carbon dioxide or CO, carbon monoxide
 (iv) Na₂CO₃, sodium carbonate (C and O forms the radical ion CO₃²⁻)
 e) Carbon dioxide is a covalent compound with a simple molecular structure. It is made up of small molecules with weak intermolecular forces of attraction between molecules. Little amount of heat energy is needed to overcome these weak forces. Thus, it has a low boiling point and is a gas at r.t.p.

Notes: I. Type of Bonding - Ionic & Covalent



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Table of Contents Prentice Hall Answer Key 7th Grade Worksheets

1. Understanding the eBook Prentice Hall Answer Key 7th Grade Worksheets
 - The Rise of Digital Reading Prentice Hall Answer Key 7th Grade Worksheets
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Answer Key 7th Grade Worksheets
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Answer Key 7th Grade Worksheets
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Answer Key 7th Grade Worksheets
 - Personalized Recommendations
 - Prentice Hall Answer Key 7th Grade Worksheets User Reviews and Ratings
 - Prentice Hall Answer Key 7th Grade Worksheets and Bestseller Lists

5. Accessing Prentice Hall Answer Key 7th Grade Worksheets Free and Paid eBooks
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 - Prentice Hall Answer Key 7th Grade Worksheets eBook Subscription Services
 - Prentice Hall Answer Key 7th Grade Worksheets Budget-Friendly Options
6. Navigating Prentice Hall Answer Key 7th Grade Worksheets eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Answer Key 7th Grade Worksheets Compatibility with Devices
 - Prentice Hall Answer Key 7th Grade Worksheets Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Answer Key 7th Grade Worksheets
 - Highlighting and Note-Taking Prentice Hall Answer Key 7th Grade Worksheets
 - Interactive Elements Prentice Hall Answer Key 7th Grade Worksheets
8. Staying Engaged with Prentice Hall Answer Key 7th Grade Worksheets
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Answer Key 7th Grade Worksheets
9. Balancing eBooks and Physical Books Prentice Hall Answer Key 7th Grade Worksheets
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Answer Key 7th Grade Worksheets
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Answer Key 7th Grade Worksheets
 - Setting Reading Goals Prentice Hall Answer Key 7th Grade Worksheets
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Answer Key 7th Grade Worksheets
 - Fact-Checking eBook Content of Prentice Hall Answer Key 7th Grade Worksheets
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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