

Chapter 7 Chemical Reactions

Section 7.2 Types of Reactions

(pages 199–205)

This section discusses how chemical reactions are classified into different types.

Reading Strategy (page 199)

Previewing. Skim the section and begin a concept map like the one below that identifies types of reactions with a general form.

As you read, add the general forms of each type of reaction. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.



Classifying Reactions (pages 199–204)

1. Name five general types of chemical reactions. Synthesis, decomposition, single replacement, double replacement, and combustion.
2. Circle the letter of each equation that represents a synthesis reaction.
 - a. $2\text{Na} + \text{Cl}_2 \longrightarrow 2\text{NaCl}$
 - b. $2\text{NaCl} \longrightarrow 2\text{Na} + \text{Cl}_2$
 - c. $2\text{H}_2\text{O} \longrightarrow 2\text{H}_2 + \text{O}_2$
 - d. $2\text{H}_2 + \text{O}_2 \longrightarrow 2\text{H}_2\text{O}$
3. Is the following sentence true or false? A decomposition reaction is the opposite of a synthesis reaction. true
4. Write the equation for the decomposition of calcium carbonate into calcium oxide and carbon dioxide. $\text{CaCO}_3 \longrightarrow \text{CaO} + \text{CO}_2$
5. Circle the letter of the correct answer. Copper reacts with silver nitrate in a single-replacement reaction. What are the products of this reaction?
 - a. copper(II) nitride and silver oxide
 - b. copper(II) nitrate and silver
 - c. copper(II) oxide and silver nitrate
 - d. copper, nitrogen, and silver oxide

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O. Jerry Parker, Gary L. Breneman



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physical chemistry with mathematical rigor and conceptual clarity and develops the modern biological applications alongside the physical principles The traditional presentations of physical chemistry are augmented with material that makes these chemical ideas biologically relevant applying physical principles to the understanding of the complex problems of 21st century biology

KEY TOPICS Physical Chemistry Biology MARKET For all readers interested in physical chemistry and biology

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Whitaker's Book List, 1991 Chemical Equations: Chemistry Essentials Practice Workbook with 200+ Reactions to Balance Jake Mariska, 2021-05-08 A chemical equation is something you will encounter every day in chemistry It is a written representation using numbers and symbols of the process that occurs during a chemical reaction Master the art of balancing chemical reactions through examples and practice 10 examples are fully solved step by step with explanations to serve as a guide Over 200 chemical equations provide ample practice Exercises start out easy and grow progressively more challenging and involved Answers to every problem are tabulated at the back of the book A chapter of pre balancing exercises helps develop essential counting skills Opening chapter reviews pertinent concepts and ideas

Balancing Chemical Equations Worksheets (Over 200 Reactions to Balance) Chris McMullen, 2016-01-12 Master the art of balancing chemical reactions through examples and practice 10 examples are fully solved step by step with explanations to serve as a guide Over 200 chemical equations provide ample practice Exercises start out easy and grow progressively more challenging and involved Answers to every problem are tabulated at the back of the book A chapter of pre balancing exercises helps develop essential counting skills Opening chapter reviews pertinent concepts and ideas Not just for students Anyone who enjoys math and science puzzles can enjoy the challenge of balancing these chemical reactions

Prentice Hall Science Explorer Chemical Interactions Adapted Reading and Study Workbook Michael J. Padilla, Ioannis Miaoulis, Martha Cyr, Prentice Hall Staff, 2005-05-01 1 Atoms and Bonding 2 Chemical Reactions 3 Acids Bases and Solutions 4 Carbon Chemistry

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Writing Chemical & Equations Marcia L. Gillette, 1993-01-01

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Table of Contents Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets

1. Understanding the eBook Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - The Rise of Digital Reading Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Personalized Recommendations
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets User Reviews and Ratings
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets and Bestseller Lists
5. Accessing Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets Free and Paid eBooks
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets Public Domain eBooks
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets eBook Subscription Services
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets Budget-Friendly Options
6. Navigating Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets Compatibility with Devices
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Highlighting and Note-Taking Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Interactive Elements Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
8. Staying Engaged with Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
- Setting Reading Goals Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
- Fact-Checking eBook Content of Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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