

## Chapter 7 Chemical Reactions

## Section 7.2 Types of Reactions

(pages 199–205)

This section discusses how chemical reactions are classified into different types.

## Reading Strategy (page 199)

**Previewing.** Skim the section and begin a concept map like the one below that identifies types of reactions with a general form.

As you read, add the general forms of each type of reaction. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.



## Classifying Reactions (pages 199–204)

1. Name five general types of chemical reactions. Synthesis, decomposition, single replacement, double replacement, and combustion.
2. Circle the letter of each equation that represents a synthesis reaction.
  - a.  $2\text{Na} + \text{Cl}_2 \longrightarrow 2\text{NaCl}$
  - b.  $2\text{NaCl} \longrightarrow 2\text{Na} + \text{Cl}_2$
  - c.  $2\text{H}_2\text{O} \longrightarrow 2\text{H}_2 + \text{O}_2$
  - d.  $2\text{H}_2 + \text{O}_2 \longrightarrow 2\text{H}_2\text{O}$
3. Is the following sentence true or false? A decomposition reaction is the opposite of a synthesis reaction. true
4. Write the equation for the decomposition of calcium carbonate into calcium oxide and carbon dioxide.  $\text{CaCO}_3 \longrightarrow \text{CaO} + \text{CO}_2$
5. Circle the letter of the correct answer. Copper reacts with silver nitrate in a single-replacement reaction. What are the products of this reaction?
  - a. copper(II) nitride and silver oxide
  - b. copper(II) nitrate and silver
  - c. copper(II) oxide and silver nitrate
  - d. copper, nitrogen, and silver oxide

# Prentice Hall Chemistry Section 17 Review Answers

**Tim D. White, Pieter A. Folken**



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