# 12

### STOICHIOMETRY

#### SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353-358)

This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

#### Using Everyday Equations (pages 353-355)

1. How can you determine the quantities of reactants and products in a chemical reaction?

You can use the balanced equation.

- Quantity usually means the \_\_\_\_\_\_\_ of a substance expressed in grams or moles.
- A bookcase is to be built from 3 shelves (Sh), 2 side boards (Sb), 1 top (T), 1 base (B), and 4 legs (L). Write a "balanced equation" for the construction of this bookcase.

3Sh + 2Sb + T + B + 4L = Sh.Sb.TBL.

#### ▶ Using Balanced Chemical Equations (page 354)

- Is the following sentence true or false? Stoichiometry is the calculation of quantities in chemical reactions.
- 5. Calculations using balanced equations are called sloichiometric calculations

#### ► Interpreting Chemical Equations (pages 356-357)

6. From what elements is ammonia produced? How is it used?

Ammonia molecules are composed of nitrogen and hydrogen; it is used as a fertilizer.

Circle the letter of the term that tells what kind of information you CANNOT get from a chemical equation.

a. moles

d. volume

b. mass

e. number of particles

(c) size of particles

## **Prentice Hall Chemistry Stoichiometry Test**

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#### **Prentice Hall Chemistry Stoichiometry Test:**

**Prentice Hall Chemistry** Antony C. Wilbraham, 2008 Prentice Hall Chemistry .2000 Stoichiometry and Thermodynamics of Metallurgical Processes Y. K. Rao, 1985-10-31 Originally published in 1985 this textbook provides a thorough and comprehensive coverage of a wide range of topics in stoichiometry and thermodynamics with special emphasis on applications to metallurgical processes This book will be welcomed as a text for courses in elementary and advanced thermodynamics and stoichiometry Treatise on Analytical Chemistry: Theory and practice. v Izaak Maurits 10th European Conference on Games Based Learning, Kolthoff, Philip J. Elving, 1959 The Software Encyclopedia 2001 Sourcebook for Chemistry and Physics David R. Hittle,1973 Three Cognitive Skills in Chemistry and Their ,2001 **Application to Stoichiometry** Ardas Ozsogomonyan, 1977 **Solving Problems in Chemistry** Rod O'Connor, Charles Mickey, Alton Hassell, 1977 Intermetallic Compounds: Practice Jack Hall Westbrook, Robert Louis Fleischer, 1995 This two volume set provides a single source for scientists and engineers interested in intermetallics The work consists of nearly 80 chapters covering fundamental theory experimental aspects practical applications present and potential and critical **Vocational-technical Learning Materials** Bruce Reinhart, 1974 assessment American Book Publishing Record Cumulative, 1950-1977 R.R. Bowker Company. Department of Bibliography, 1978 Sm Fundamentals Chemistry T/I Burns.1999-03 *Films and Other Materials for Projection* Library of Congress,1978 **Problems and Problem Solving** in Chemistry Education Georgios Tsaparlis, 2021-05-17 Problem solving is central to the teaching and learning of chemistry at secondary tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data looking for patterns and making deductions As an important higher order thinking skill problem solving also constitutes a major research field in science education Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving The following situations are considered some general others with a focus on specific areas of chemistry quantitative problems qualitative reasoning metacognition and resource activation deconstructing the problem solving process an overview of the working memory hypothesis reasoning with the electron pushing formalism scaffolding organic synthesis skills spectroscopy for structural characterization in organic chemistry enzyme kinetics problem solving in the academic chemistry laboratory chemistry problem solving in context team based active learning technology for molecular representations IR spectra simulation and computational quantum chemistry tools The book concludes with methodological and epistemological issues in problem solving research and other perspectives in problem solving in chemistry With a foreword by George Bodner **Innovative Curriculum Materials** ,1999 **Chemistry** Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02 Experimental Biochemistry J. Stenesh, 1984 Ground-water Geochemistry at the South Ramp Jet Fuel Leak, Diego Garcia Atoll, 1993-94 Gordon W. Tribble, 2001

 $\textbf{Water-resources Investigations Report}\ , 2001$ 

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  - Chapter 4: Prentice Hall Chemistry Stoichiometry Test in Specific Contexts
  - ∘ Chapter 5: Conclusion
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