

11

CHEMICAL REACTIONS

Practice Problems

In your notebook, solve the following problems. Use the 3-step problem-solving approach you learned in Chapter 1.

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS

1. Write the skeleton equation for the reaction between hydrogen and oxygen that produces water.
2. Write the skeleton equation for the reaction that produces iron(II) sulfide from iron and sulfur.
3. Write the skeleton equation representing the heating of magnesium carbonate to produce solid magnesium oxide and carbon dioxide gas.
4. Write a balanced equation for the production of HCl gas from its elements.
5. Write a sentence that completely describes the chemical reaction represented by this balanced equation.



6. Write the word equation for the following equation. Write a sentence fully describing the reaction. Is the equation correctly balanced? Explain.



7. Write a balanced equation representing the formation of aqueous sulfuric acid from water and sulfur trioxide gas.
8. Write a balanced equation from this word equation.
aqueous silver nitrate + copper metal \rightarrow silver metal + aqueous copper nitrate
9. Write a balanced equation for the following word equation.



SECTION 11.2 TYPES OF CHEMICAL REACTIONS

1. Write a balanced equation representing the reaction of magnesium with oxygen gas to produce magnesium oxide.
2. Write the balanced equation for the reaction that occurs between aluminum and fluorine.
3. Write the balanced equation for the production of oxygen gas and potassium chloride from the decomposition of potassium chlorate.
4. Write the balanced equation for the reaction between hydrochloric acid and calcium metal. The products are hydrogen gas and calcium chloride.
5. Write the balanced equation for the combustion of propane (C_3H_8) to produce carbon dioxide and water vapor.
6. Write the balanced equation for the reaction between iron(III) chloride and sodium hydroxide. The products are iron(III) hydroxide and sodium chloride.

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1. Understanding the eBook Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
 - The Rise of Digital Reading Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets

- Personalized Recommendations
 - Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets User Reviews and Ratings
 - Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets and Bestseller Lists
5. Accessing Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets Free and Paid eBooks
 - Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets Public Domain eBooks
 - Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets eBook Subscription Services
 - Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets Budget-Friendly Options
 6. Navigating Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets Compatibility with Devices
 - Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
 - Highlighting and Note-Taking Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
 - Interactive Elements Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
 8. Staying Engaged with Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
 9. Balancing eBooks and Physical Books Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
 - Setting Reading Goals Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets

- Fact-Checking eBook Content of Prentice Hall Chemistry Chapter 11 Chemical Reactions Worksheets
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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