

## 11-3 Practice Problems

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1. Identify the limiting reactant when 1.22-g of  $O_2$  reacts with 1.65-g of  $H_2$  to produce water.

2. Identify the limiting reactant when 4.68-g of Fe reacts with 2.88-g of S to produce  $FeS$ .

3. Identify the limiting reactant when 2.87-g of  $Mg(OH)_2$  reacts with 12.84-g of  $HCl$  to form  $MgCl_2$  and water.

4. Identify the limiting reactant when 6.25-g of  $AgNO_3$  reacts with 4.12-g of  $NaCl$  to form  $NaNO_3$  and  $AgCl$ .

5. Identify the limiting reactant when 7.81-g of  $HCl$  reacts with 5.24-g of  $NaOH$  to produce  $NaCl$  and  $H_2O$ .

6. Identify the limiting reactant when 6.33-g of  $H_2SO_4$  reacts with 5.92-g of  $NaOH$  to produce  $Na_2SO_4$  and water.

7. Identify the limiting reactant when 49.25-g of  $CaC_2$  reacts with 33.71-g of water to produce  $Ca(OH)_2$  and  $C_2H_2$ .

8. Identify the limiting reactant when 63.14-g of  $CaCl_2$  reacts with 74.68-g of  $Na_2CO_3$  to produce  $CaCO_3$  and  $NaCl$ .

9. Identify the limiting reactant when 4.687-g of  $SF_4$  reacts with 6.281-g of  $LiO_2$  to produce  $IF_3$  and  $SiO_2$ .

10. If 4.1-g of Cr is heated with 9.3-g of  $Cl_2$ , what mass  $CrCl_3$  will be produced?

11. What mass of  $SO_2$  is produced from the reaction between 31.5-g of  $S_8$  and 8.65-g of  $O_2$ ?

12. What mass of  $SO_2$  is produced from the reaction of 12.4-g of  $SO_2$  and 3.45-g of  $O_2$ ?

13. What mass of  $H_2SO_4$  is produced from the reaction of 6.58-g of  $SO_3$  and 1.64-g of  $H_2O$ ?

14. What mass of  $CdS$  is produced if 8.47-g of cadmium reacts with 2.51-g of Sulfur?

# Prentice Hall Chemistry Chapter 3 Practice Problems Answers

**Branko N. Popov**



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