

## Chapter 7 Chemical Reactions

**Section 7.1 Describing Reactions**

(pages 192–198)

*This section discusses the use of chemical equations and how to balance them. It also demonstrates the use of calculations in chemistry.*

**Reading Strategy (page 192)**

**Monitoring Your Understanding** Preview the Key Concepts, topic headings, vocabulary, and figures in this section. List two things you expect to learn. After reading, state what you learned about each item you listed. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Expect to Learn	What I Learned
Answers may vary. Possible answers: How to balance chemical equations	Answers may vary. Possible answers: An unbalanced equation can be balanced by changing the coefficients.
How to convert from mass to moles	The mass of a substance can be converted to moles by using the molar mass as a conversion factor.

**Chemical Equations (pages 192–193)**

- Is the following sentence true or false? The new substances formed as a result of a chemical reaction are called products. true
- Circle the letter of each sentence that is a correct interpretation of the chemical equation  $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$ .
  - Carbon and oxygen react and form carbon monoxide.
  - ☒ Carbon and oxygen react and form carbon dioxide.
  - Carbon dioxide yields carbon and oxygen.
  - ☒ The reaction of carbon and oxygen yields carbon dioxide.
- Is the following sentence true or false? The law of conservation of mass states that mass is neither created nor destroyed in a chemical reaction. true
- Circle the letter of the correct answer. According to the equation  $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$ , how many carbon atoms react with 14 molecules of oxygen to form 14 molecules of carbon dioxide?
  - 1
  - 7
  - ☒ 14
  - 28
- In the reaction represented by the equation  $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$ , the mass of carbon dioxide produced equals the total mass of carbon and oxygen that reacted.

# Prentice Hall Chemistry Workbook Answers Ch 7

**Patrick Vollmar**



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### **Table of Contents Prentice Hall Chemistry Workbook Answers Ch 7**

1. Understanding the eBook Prentice Hall Chemistry Workbook Answers Ch 7
  - The Rise of Digital Reading Prentice Hall Chemistry Workbook Answers Ch 7
  - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Workbook Answers Ch 7
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Prentice Hall Chemistry Workbook Answers Ch 7
  - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Workbook Answers Ch 7
  - Personalized Recommendations
  - Prentice Hall Chemistry Workbook Answers Ch 7 User Reviews and Ratings
  - Prentice Hall Chemistry Workbook Answers Ch 7 and Bestseller Lists

5. Accessing Prentice Hall Chemistry Workbook Answers Ch 7 Free and Paid eBooks
  - Prentice Hall Chemistry Workbook Answers Ch 7 Public Domain eBooks
  - Prentice Hall Chemistry Workbook Answers Ch 7 eBook Subscription Services
  - Prentice Hall Chemistry Workbook Answers Ch 7 Budget-Friendly Options
6. Navigating Prentice Hall Chemistry Workbook Answers Ch 7 eBook Formats
  - ePub, PDF, MOBI, and More
  - Prentice Hall Chemistry Workbook Answers Ch 7 Compatibility with Devices
  - Prentice Hall Chemistry Workbook Answers Ch 7 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Workbook Answers Ch 7
  - Highlighting and Note-Taking Prentice Hall Chemistry Workbook Answers Ch 7
  - Interactive Elements Prentice Hall Chemistry Workbook Answers Ch 7
8. Staying Engaged with Prentice Hall Chemistry Workbook Answers Ch 7
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Prentice Hall Chemistry Workbook Answers Ch 7
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Workbook Answers Ch 7
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Prentice Hall Chemistry Workbook Answers Ch 7
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Workbook Answers Ch 7
  - Setting Reading Goals Prentice Hall Chemistry Workbook Answers Ch 7
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Workbook Answers Ch 7
  - Fact-Checking eBook Content of Prentice Hall Chemistry Workbook Answers Ch 7
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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