

## 12

## STOICHIOMETRY

**SECTION 12.1 THE ARITHMETIC OF EQUATIONS** (pages 353–358)

This section explains how to calculate the masses of reactants required or products formed in a stoichiometric process. It teaches you how to interpret chemical equations in terms of reacting moles, representative particles, masses, and gas volume at STP.

**► Using Everyday Equations** (pages 353–354)

1. How can you determine the quantities of reactants and products in a chemical reaction?

You must enter **2** or **balancing** or **equation**.

2. Quantity usually means the \_\_\_\_\_ **amount** \_\_\_\_\_ of a substance expressed in grams or moles.

3. A bookkeeper is on the fault lines 3 shelves (SF), 2 side boards (SB), 1 top (T), 1 base (B), and 4 legs (L). Write a “balanced equation” for the construction of this bookcase.

**$3\text{SF} + 2\text{SB} + 1\text{T} + 1\text{B} + 4\text{L} \rightarrow 1\text{bookcase}$**

**► Using Balanced Chemical Equations** (page 354)

4. In the following sentence (use an Internet Search engine to do calculation of quantities in chemical reactions), \_\_\_\_\_ **is not** \_\_\_\_\_.

5. Calculations using balanced equations are called \_\_\_\_\_ **stoichiometric calculations** \_\_\_\_\_.

**► Interpreting Chemical Equations** (pages 354–357)

6. From what elements is ammonia produced? How is it used?

**Nitrogen and hydrogen are combined at nitrogen’s high temperature. It is used as a fertilizer.**

7. Circle the letter of the term that tells what kind of information you **CANNOT** get from a chemical equation.

- |                             |                        |
|-----------------------------|------------------------|
| a. moles                    | d. volume              |
| b. mass                     | e. number of particles |
| <b>c. size of particles</b> |                        |

# The Arithmetic Of Equations 12 Answers

**Samuel Maynard**



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