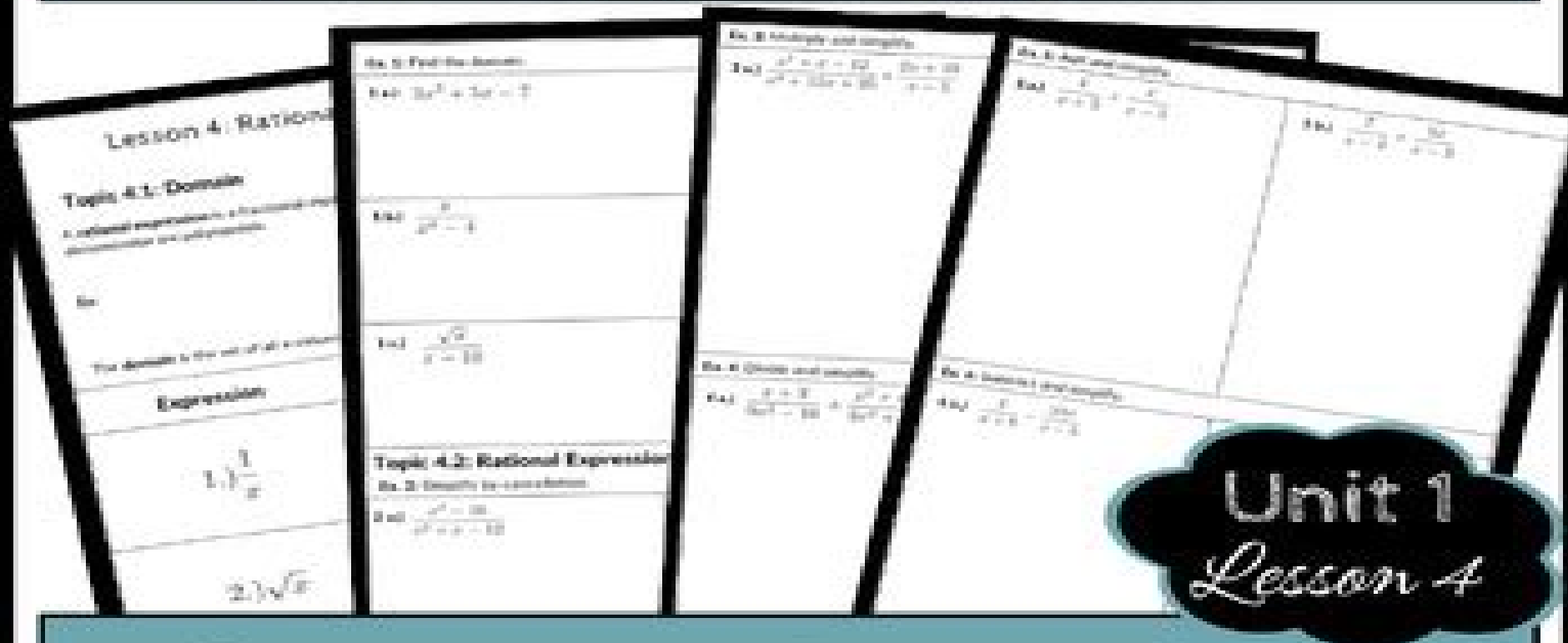


# PRECALCULUS

## Unit 1: Fundamentals

### Lesson 4: Rational Expressions



Guided notes & Problem Set

# Precalculus Unit Lesson 01

**Charlene E. Beckmann, Denisse R.  
Thompson, Rheta N. Rubenstein**



## Precalculus Unit Lesson 01:

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Precalculus The Easy Way provides easy to follow lessons with comprehensive review and practice This edition features a brand new design and new content structure with illustrations and practice questions An essential resource for High school and college courses Virtual learning Learning pods Homeschooling Precalculus The Easy Way covers Algebraic Methods Functions and Graphs Complex Numbers Polynomial and Rational Functions Calculus Preview And more *Eureka Math Precalculus Study Guide* Great Minds, 2016-07-14 The team of teachers and mathematicians who created Eureka Math believe that it s not enough for students to know the process for solving a problem they need to know why that process works That s why students who learn math with Eureka can solve real world problems even those they have never encountered before The Study Guides are a companion to the Eureka Math program whether you use it online or in print The guides collect the key components of the curriculum for each grade in a single volume They also unpack the standards in detail so that anyone even non Eureka users can benefit The guides are particularly helpful for teachers or trainers seeking to undertake or lead a meaningful study of the grade level content in a way that highlights the coherence between modules and topics We re here to make sure you succeed with an ever growing library of resources Take advantage of the full set of Study Guides available for each grade PK 12 or materials at eureka math org such as free implementation and pacing guides material lists parent resources and more

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are to help all students become mathematically proficient Throughout this book you are encouraged to work in cooperative teams This strategy is designed to help you develop a mathematics learning community and build a professional network that will be a valuable resource during your professional career Hopefully you will experience the benefits of engaging in rich mathematical discussions with peers and consider how to encourage such learning environments in your own classrooms Lesson planning is another element pervasive throughout this text To help teachers plan for effective student centered lessons the Question Response Support QRS Guide is introduced in Lesson 1 1 and used throughout the remainder of the lessons The QRS Guide is a tool on which teachers may record tasks or questions Q for students expected and observed student responses R and teacher support S in the form of additional just enough questions to support students in their progress on the task In each unit teachers expand their repertoire of teaching and learning elements and strategies and incorporate these elements as they plan additional lesson segments In Unit 4 lesson planning is formally introduced as teachers put together elements from previous units into complete cohesive lesson plans

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web jul 3 2019 a covalent bond in chemistry is a chemical link between two atoms or ions in which the electron pairs are shared between them a covalent bond may also be termed a molecular bond covalent bonds form between two nonmetal atoms with identical or relatively close electronegativity values this type of bond may also be found in other

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web important key terms to practice from chapter 6 terms in this set 20 covalent bond a bond formed when atoms share one or more pairs of electrons molecular orbital the region of high probability of finding an electron that is being shared between two nuclei in a covalent bond bond length

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web the carbon hydrogen chains in vulcanized rubber are held together by two sulfur atoms that form covalent bonds between the chains these covalent bonds are commonly called disulfide bridges explore other molecules that have such disulfide bridges present your findings to the class

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web because of the fact that polar covalent bonds have higher melting and boiling points why do dipole forces have an impact on the melting and boiling points of substances this is because the dipole forces are an added force of attraction that slows the molecules down making it take more energy to melt or boil them