#### Precipitation Reactions Worksheet Key

For each of the following reactants, predict whether a precipitation reaction will take place between them. If there is no reaction, write "no reaction". If there is a reaction, write the complete, complete ionic, and net ionic equations that describe the reaction.

- $\begin{array}{lll} \text{2. Li}_2\text{CO}_3(aq) + \text{Co}(\text{C}_2\text{H}_3\text{O}_2)_2(aq) & \rightarrow & 2\text{Li}\text{C}_2\text{H}_3\text{O}_2(aq) + \text{Co}\text{CO}_3(s) \\ 2\text{Li}^*(aq) + \text{CO}_3^{2^-}(aq) + \text{Co}^{2^+}(aq) + 2\text{C}_2\text{H}_3\text{O}_2^-(aq) & \rightarrow & 2\text{Li}^*(aq) + 2\text{C}_2\text{H}_3\text{O}_2^-(aq) + \text{Co}\text{CO}_3(s) \\ \text{CO}_3^{2^-}(aq) + \text{Co}^{2^+}(aq) & \rightarrow & \text{Co}\text{CO}_3(s) \\ \end{array}$
- b.  $2\text{Fe}(\text{NO}_3)_3(aq) + 3\text{K}_2\text{S}(aq) \rightarrow \text{Fe}_2\text{S}_3(s) + 6\text{KNO}_3(aq)$   $2\text{Fe}^{3+}(aq) + 6\text{NO}_3^-(aq) + 6\text{K}^+(aq) + 3\text{S}^{2-}(aq) \rightarrow \text{Fe}_2\text{S}_3(s) + 6\text{K}^+(aq) + 6\text{NO}_3^-(aq)$  $2\text{Fe}^{3+}(aq) + 3\text{S}^{2-}(aq) \rightarrow \text{Fe}_2\text{S}_3(s)$
- c.  $Pb(NO_3)_2(aq) + Li_2SO_4(aq) \rightarrow PbSO_4(s) + 2LiNO_3(aq)$   $Pb^{2+}(aq) + 2NO_3^-(aq) + 2Li^+(aq) + SO_4^{2-}(aq) \rightarrow PbSO_4(s) + 2Li^+(aq) + 2NO_3^-(aq)$  $Pb^{2+}(aq) + SO_4^{2-}(aq) \rightarrow PbSO_4(s)$
- d. NH<sub>4</sub>Cl(aq) + Cu(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub>(aq) No reaction

#### CHEMISTRY 151 - PRECIPITATION REACTIONS KEY

For each of the following reactants, predict whether a precipitation reaction will take place between them. If there is no reaction, write "no reaction". If there is a reaction, write the complete, complete ionic, and net ionic equations that describe the reaction.

- b.  $K_2S(aq) + ZnCl_2(aq) \rightarrow 2KCl(aq) + ZnS(s)$   $2K^+(aq) + S^{2-}(aq) + Zn^{2+}(aq) + 2Cl^-(aq) \rightarrow 2K^+(aq) + 2Cl^-(aq) + ZnS(s)$  $S^{2-}(aq) + Zn^{2+}(aq) \rightarrow ZnS(s)$
- c.  $2NH_4Br(aq) + Pb(C_2H_3O_2)_2(aq) \rightarrow 2NH_4C_2H_3O_2(aq) + PbBr_2(s)$   $2NH_4^+(aq) + 2Br^-(aq) + Pb^{2+}(aq) + 2C_2H_3O_2^-(aq)$   $\rightarrow 2NH_4^+(aq) + 2C_2H_3O_2^-(aq) + PbBr_2(s)$  $2Br^-(aq) + Pb^{2+}(aq) \rightarrow PbBr_2(s)$
- d.  $Na_2SO_4(aq) + Ba(NO_3)_2(aq) \rightarrow 2NaNO_3(aq) + BaSO_4(s)$   $2Na^*(aq) + SO_4^{2-}(aq) + Ba^{2*}(aq) + 2NO_3^{-}(aq)$   $\rightarrow 2Na^*(aq) + 2NO_3^{-}(aq) + BaSO_4(s)$  $SO_4^{2-}(aq) + Ba^{2*}(aq) \rightarrow BaSO_4(s)$
- e. K<sub>3</sub>PO<sub>4</sub>(aq) + (NH<sub>4</sub>)<sub>2</sub>CO<sub>3</sub>(aq) No Reaction

# **Precipitation Reactions Answer Sheet**

# **Akash Tiwari**

#### **Precipitation Reactions Answer Sheet:**

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This book delves into Precipitation Reactions Answer Sheet. Precipitation Reactions Answer Sheet is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Precipitation Reactions Answer Sheet, encompassing both the fundamentals and more intricate discussions.

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- 6. In chapter 5, the author will draw a conclusion about Precipitation Reactions Answer Sheet. This chapter will summarize the key points that have been discussed throughout the book.
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