

**Chemistry Chapter 10 Test Practice Problems**

Convert each of the following to **moles**:

- a. 22.0 g  $\text{NO}_2$
- b. 190 g  $\text{MgF}_2$
- c.  $2.35 \times 10^{25}$  atom K
- d. 18 molecules  $\text{Na}_2\text{O}$

Convert each of the following to **grams**:

- a. 4.10 mol  $\text{CaCO}_3$
- b. 10.2 mol  $\text{C}_2\text{P}_3$
- c.  $3.01 \times 10^{23}$  atoms Li
- d.  $1.2 \times 10^{18}$  molecules  $\text{C}_3\text{H}_{10}\text{O}_5$

Convert each of the following to **liters**:

- a. 1.567 mol Ne
- b.  $3.56 \times 10^{27}$  atoms of He
- c. 6.0 g  $\text{O}_2$
- d.  $2.51 \times 10^{20}$  molecules  $\text{CO}_2$

Calculate the **percentage composition** for each element in the following compounds:

- |                    |                            |
|--------------------|----------------------------|
| a. $\text{MgCl}_2$ | c. $\text{Pb}_2\text{O}_3$ |
| b. $\text{CaI}_2$  | d. $\text{AuSO}_4$         |

# Prentice Hall Chemistry Answer Key Chemical Quantities

**Richard Bailey**



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