

### Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility**: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.  
**saturated solution**: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution**: a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible** describes liquids that dissolve in each other. **immiscible** describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No; if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47.  $5.55 \times 10^2$  g  $\text{AgNO}_3$
48. Solubility increases with pressure.
49. a.  $1.6 \times 10^{-2}$  g/L  
b.  $4.7 \times 10^{-2}$  g/L
50. Dilute and concentrated are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.  
a. 1.3M KCl  
b.  $5.3 \times 10^{-2}$  M  $\text{MgCl}_2$
52.  $2.00 \times 10^2$  mL
53. a.  $5.0 \times 10^{-2}$  mol NaCl, 20 g NaCl  
b. 1.0 mol  $\text{KNO}_3$ ,  $1.0 \times 10^2$  g  $\text{KNO}_3$   
c.  $2.5 \times 10^2$  mol  $\text{CaCl}_2$ , 2.0 g  $\text{CaCl}_2$
54. a.  $2.3 \times 10^2$  g NaCl  
b. 2.0 g  $\text{MgCl}_2$
55. a. 10% (v/v) ethanol  
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water  
b. 1.5M  $\text{KNO}_3$   
c. 0.100M  $\text{MgCl}_2$
58. The effective molality of the  $\text{Ca}(\text{NO}_3)_2$  solution is 3m. The effective molality of the  $\text{NaNO}_3$  solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g  $\text{H}_2\text{O}$  to 32.0 g  $\text{C}_2\text{H}_5\text{OH}$ .
63. a.  $-100.26^\circ\text{C}$   
b.  $-101.54^\circ\text{C}$
64. a.  $-4.48^\circ\text{C}$   
b.  $-2.2^\circ\text{C}$
65. a.  $-1.1^\circ\text{C}$   
b.  $-0.24^\circ\text{C}$   
c.  $-1.3^\circ\text{C}$

### Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

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Baker**



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  - Advantages of eBooks Over Traditional Books
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  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
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  - Popular eBook Platforms
  - Features to Look for in an Prentice Hall Chemistry Answers Ch 18
  - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Answers Ch 18
  - Personalized Recommendations
  - Prentice Hall Chemistry Answers Ch 18 User Reviews and Ratings
  - Prentice Hall Chemistry Answers Ch 18 and Bestseller Lists

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  - Prentice Hall Chemistry Answers Ch 18 Budget-Friendly Options
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  - ePub, PDF, MOBI, and More
  - Prentice Hall Chemistry Answers Ch 18 Compatibility with Devices
  - Prentice Hall Chemistry Answers Ch 18 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Answers Ch 18
  - Highlighting and Note-Taking Prentice Hall Chemistry Answers Ch 18
  - Interactive Elements Prentice Hall Chemistry Answers Ch 18
8. Staying Engaged with Prentice Hall Chemistry Answers Ch 18
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Prentice Hall Chemistry Answers Ch 18
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Answers Ch 18
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Prentice Hall Chemistry Answers Ch 18
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Answers Ch 18
  - Setting Reading Goals Prentice Hall Chemistry Answers Ch 18
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Answers Ch 18
  - Fact-Checking eBook Content of Prentice Hall Chemistry Answers Ch 18
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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