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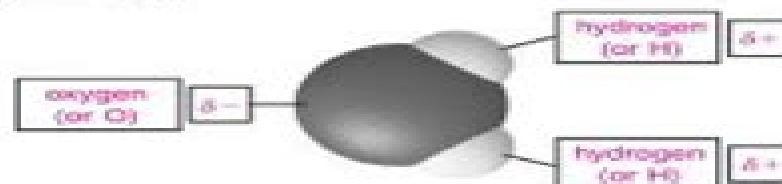
WATER AND AQUEOUS SYSTEMS

SECTION 15.1 WATER AND ITS PROPERTIES (pages 445–449)

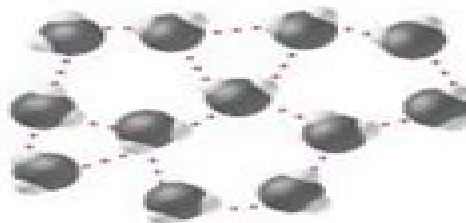
This section describes the properties of water in the liquid and solid states and explains how hydrogen bonding affects the surface tension and vapor pressure of water.

► Water in the Liquid State (pages 445–447)

1. What unique substance is essential to all life on Earth?
water
2. Approximately what fraction of Earth's surface is covered in water? 3/4
3. Circle the letter next to each sentence that is true concerning water molecules.
 - a. Each O — H covalent bond in a water molecule is nonpolar.
 - b.** In a water molecule, the less electronegative hydrogen atoms acquire a partial positive charge and the oxygen atom acquires a partial negative charge.
 - c. Because the water molecule has an H — O — H bond angle of 105° , the molecule as a whole is polar.
4. The diagram below depicts a water molecule. Complete the labels showing the locations of the hydrogen atoms, the oxygen atom, and the regions of positive and negative charge.



5. The diagram below depicts a collection of water molecules. Draw dotted lines showing where hydrogen bonding occurs.



Prentice Hall Chemistry Answers Ch15

**Janick F. Artiola, Ian L. Pepper, Mark L.
Brusseu**



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