

13.1 THE NATURE OF GASES

P.1

Section Review

Objectives

- Describe the assumptions of the kinetic theory as it applies to gases
- Interpret gas pressure in terms of kinetic theory
- Define the relationship between Kelvin temperature and average kinetic energy

Vocabulary

- | | | |
|------------------|------------------------|-----------------------------|
| • kinetic energy | • vacuum | • pascal (Pa) |
| • kinetic theory | • atmospheric pressure | • standard atmosphere (atm) |
| • gas pressure | • barometer | |

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The kinetic theory describes the 1 of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly 2, that the particles of gas are relatively 3, move 4 of each other, and are in constant 5 motion. The 6 between particles are perfectly elastic so that the total 7 remains constant. Gas pressure results from the simultaneous collisions of billions of particles with an object. Barometers are used to measure 8 pressure. Standard conditions are defined as a temperature of 9 and a pressure of 10.

- motion
- empty space
- far apart
- independently
- random / rapid
- collisions
- Kinetic energy
- atmospheric
- 0°C / 273 K
- 1 atm / 101.3 kPa

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

ST 11. Atmospheric pressure is 760 mm Hg.

AT 12. The SI unit of pressure is the pascal.

↳ a multiple of kilo pascal

18 pts
 - 2 pt if late
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Computational Chemistry Errol G. Lewars, 2024-06-04. This is the fourth edition of the successful textbook on computational chemistry which continues to provide a comprehensive introduction to the theory and practice of computational chemistry. Notable updates include a review of references up to mid 2023 encompassing recent developments in scientific journals, books, and software. The evolving prominence of density functional theory (DFT) is emphasized and attention is given to the increasing application of artificial intelligence in computational chemistry. The book maintains key features from the previous edition, delving into the mathematical intricacies of ab initio and density functional methods at an introductory level. Clear explanations of matrix methods are provided, offering a direct approach to obtaining energy levels and molecular orbitals. Additionally, each chapter includes sets of Easier and Harder drill questions with suggested answers at the end of the book, enhancing the learning experience. The book is intended for upper year undergraduate and graduate students studying computational and theoretical chemistry, and for self study by researchers in universities and industry to whom computational chemistry may be useful.

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