

4

ATOMIC STRUCTURE

Chapter Test A

A. Matching

Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line.

Column A	Column B
<u> J </u> 1. proton	a. the total number of protons and neutrons in the nucleus of an atom
<u> G </u> 2. atom	b. the weighted average mass of the atoms in a naturally occurring sample of an element
<u> A </u> 3. mass number	c. $\frac{1}{12}$ the mass of a carbon-12 atom
<u> C </u> 4. atomic mass unit	d. the number of protons in the nucleus of an element
<u> F </u> 5. electron	e. atoms with the same number of protons but different numbers of neutrons
<u> E </u> 6. isotopes	f. negatively charged subatomic particle
<u> D </u> 7. atomic number	g. the smallest particle of an element that retains its identity in a chemical reaction
<u> B </u> 8. atomic mass	h. a horizontal row of the periodic table
<u> H </u> 9. period	i. subatomic particle with no charge
<u> I </u> 10. neutron	j. positively charged subatomic particle

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 11. Which of the following is *not* a part of Dalton's atomic theory?
- All elements are composed of atoms.
 - Atoms of the same element are alike.
 - Atoms are always in motion.
 - Atoms that combine do so in simple whole-number ratios.
- _____ 12. The nucleus of an atom is
- negatively charged and has a low density.
 - negatively charged and has a high density.
 - positively charged and has a low density.
 - positively charged and has a high density.

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Brendan G. Carr



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principles Additional overview problems have been added to the end of chapter problem sets The descriptive chemistry has been updated with many new results from the literature being included Chapter 4 Bonding in polyatomic molecules has been rewritten with greater emphasis on the use of group theory for the derivation of ligand group orbitals and orbital symmetry labels There is more coverage of supercritical fluids and green chemistry The new full colour text design enhances the presentation of the many molecular structures and 3 D images Supporting this edition Companion website featuring multiple choice questions and rotatable 3 D molecular structures available at www.reasoned.co.uk/housecroft For full information including details of lecturer material see the Contents list inside the book A Solutions Manual written by Catherine E Housecroft with detailed solutions to all end of chapter problems within the text is available for purchase separately ISBN 0131 39926 8 Catherine E Housecroft is Professor of Chemistry at the University of Basel Switzerland She is the author of a number of textbooks and has extensive teaching experience in the UK Switzerland South Africa and the USA Alan G Sharpe is a Fellow of Jesus College University of Cambridge UK and has had many years of experience teaching inorganic chemistry to undergraduates

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